

Mastering Chemistry With Ravi Arora

DPP - Daily Practice Problems

Chapter-wise Sheets

Date :

Start Time :

End Time :

CHEMISTRY

CC01

SYLLABUS : Some Basic Concepts of Chemistry

Max. Marks : 180

Marking Scheme : + 4 for correct & (-1) for incorrect

Time : 60 min.

INSTRUCTIONS : This Daily Practice Problem Sheet contains 45 MCQ's. For each question only one option is correct. Darken the correct circle/ bubble in the Response Grid provided on each page.

- Given the numbers : 161 cm, 0.161 cm, 0.0161 cm. The number of significant figures for the three numbers are
(a) 3, 4 and 5 respectively (b) 3, 3 and 4 respectively
(c) 3, 3 and 3 respectively (d) 3, 4 and 4 respectively
- If the true value for an experimental result is 6.23 and the results reported by three students X, Y and Z are :
X: 6.18 and 6.28
Y: 6.20 and 6.023
Z: 6.22 and 6.24
Which of the following option is correct :
(a) X precise, Y accurate, Z precise and accurate.
(b) X precise and accurate, Y not precise, Z precise
(c) Both X & Z precise & accurate, Y not precise.
(d) Both X & Y neither precise nor accurate, Z both precise and accurate.
- Number of grams of oxygen in 32.2 g $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ is
(a) 20.8 (b) 2.24
(c) 22.4 (d) 2.08
- 3 g of an oxide of a metal is converted to chloride completely and it yielded 5 g of chloride. The equivalent weight of the metal is
(a) 3.325 (b) 33.25
(c) 12 (d) 20
- 1 cc. N_2O at NTP contains :
(a) $\frac{1.8}{224} \times 10^{22}$ atoms (b) $\frac{6.02}{22400} \times 10^{23}$ molecules
(c) $\frac{1.32}{224} \times 10^{23}$ electrons (d) All of the above

RESPONSE GRID

1. (a)(b)(c)(d)

2. (a)(b)(c)(d)

3. (a)(b)(c)(d)

4. (a)(b)(c)(d)

5. (a)(b)(c)(d)

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6. One of the following combination which illustrates the law of reciprocal proportions ?
 (a) N_2O_3, N_2O_4, N_2O_5 (b) $NaCl, NaBr, NaI$
 (c) CS_2, CO_2, SO_2 (d) PH_3, P_2O_3, P_2O_5
7. An aqueous solution of oxalic acid dihydrate contains its 6.3g in 250 mL. The volume of 0.1 N NaOH required to completely neutralize 10 mL of this solution
 (a) 4mL (b) 20mL (c) 2mL (d) 40mL
8. The density of 3M solution of sodium chloride is 1.252 g mL^{-1} . The molality of the solution will be:
 (molar mass, $NaCl = 58.5 \text{ g mol}^{-1}$)
 (a) 260m (b) 2.18m (c) 2.79m (d) 3.00m
9. The number of atoms in 0.1 mole of a triatomic gas is :
 ($N_A = 6.02 \times 10^{23} \text{ mol}^{-1}$)
 (a) 6.026×10^{22} (b) 1.806×10^{23}
 (c) 3.600×10^{23} (d) 1.800×10^{22}
10. Match the columns.
- | Column-I | Column-II |
|---|--------------------------------------|
| A. 88 g of CO_2 | I. 0.25 mole |
| B. 6.022×10^{23} molecules of H_2O | II. 2 mole |
| C. 5.6 litres of O_2 at STP | III. 1 mole |
| D. 96 g of O_2 | IV. 6.022×10^{23} molecules |
| E. 1 mol of any gas | V. 3 mole |
- (a) A – II; B – III; C – I; D – V; E – VI
 (b) A – III; B – II; C – I; D – V; E – IV
 (c) A – II; B – I; C – III; D – V; E – IV
 (d) A – II; B – III; C – I; D – IV; E – V
11. The simplest formula of a compound containing 50% of element X (atomic mass 10) and 50% of element Y (atomic mass 20) is
 (a) XY (b) XY_3 (c) X_2Y (d) X_2Y_3
12. Which one of the following is the lightest?
 (a) 0.2 mole of hydrogen gas (b) 6.023×10^{22} molecules of nitrogen
 (c) 0.1 g of silver (d) 0.1 mole of oxygen gas
13. If N_A is Avogadro's number then number of valence electrons in 4.2 g of nitride ions (N^{3-}) is
 (a) $4.2 N_A$ (b) $2.4 N_A$
 (c) $1.6 N_A$ (d) $3.2 N_A$
14. The set of numerical coefficients that balances the equation
 $K_2CrO_4 + HCl \longrightarrow K_2Cr_2O_7 + KCl + H_2O$ is
 (a) 2, 2, 1, 2, 1 (b) 2, 2, 1, 1, 1
 (c) 2, 1, 1, 2, 1 (d) 1, 1, 2, 2, 1
15. Match the columns
- | Column-I
(Number) | Column-II
(Significant figures) |
|------------------------|------------------------------------|
| A. 29900. | I. 2 |
| B. 290 | II. 1 |
| C. 1.23×1.331 | III. 4 |
| D. 20.00 | IV. 3 |
| E. $2.783 - 1$ | V. 5 |
- (a) A – III; B – II; C – V; D – I; E – IV
 (b) A – V; B – I; C – IV; D – III; E – II
 (c) A – I; B – V; C – IV; D – III; E – II
 (d) A – V; B – IV; C – III; D – II; E – I
16. The maximum number of molecules are present in
 (a) 15 L of H_2 gas at STP (b) 5 L of N_2 gas at STP
 (c) 0.5 g of H_2 gas (d) 10 g of O_2 gas
17. The number of moles of oxygen in one litre of air containing 21% oxygen by volume, under standard conditions are
 (a) 0.0093 mole (b) 0.21 mole
 (c) 2.10 mole (d) 0.186 mole
18. Assuming fully decomposed, the volume of CO_2 released at STP on heating 9.85 g of $BaCO_3$ (Atomic mass, Ba = 137) will be
 (a) 1.12 L (b) 2.24 L
 (c) 4.06 L (d) 0.84 L

RESPONSE GRID	5. (a)(b)(c)(d)	6. (a)(b)(c)(d)	7. (a)(b)(c)(d)	8. (a)(b)(c)(d)	9. (a)(b)(c)(d)
	10. (a)(b)(c)(d)	11. (a)(b)(c)(d)	12. (a)(b)(c)(d)	13. (a)(b)(c)(d)	14. (a)(b)(c)(d)
	15. (a)(b)(c)(d)	16. (a)(b)(c)(d)	17. (a)(b)(c)(d)	18. (a)(b)(c)(d)	19. (a)(b)(c)(d)

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19. The ratio of the molar amounts of H_2S needed to precipitate the metal ions from 20 mL each of 1 M $Ca(NO_3)_2$ and 0.5M $CuSO_4$ is
 (a) 1 : 1 (b) 2 : 1 (c) 1 : 2 (d) indefinite
20. Consider the following statements.
 (i) Atoms of H, O, N and C have identical properties but different mass.
 (ii) Matter is divisible into atoms which are further indivisible.
 (iii) The ratio of N: H in NH_3 is 1 : 3 and N : O in nitric oxide is 2 : 1.
 (iv) Dalton's atomic theory support law of conservation of mass.
 Which of the following pairs of statements is true according to Dalton's atomic theory ?
 (a) (i) and (ii) (b) (ii) and (iii)
 (c) (ii) and (iv) (d) (i) and (iv)
21. How many moles of $Al_2(SO_4)_3$ would be in 50 g of the substance ?
 (a) 0.083 mole (b) 0.952 mole
 (c) 0.481 mole (d) 0.140 mole
22. Experimentally it was found that a metal oxide has formula $M_{0.98}O$. Metal M, present as M^{2+} and M^{3+} in its oxide. Fraction of the metal which exists as M^{3+} would be :
 (a) 7.01% (b) 4.08% (c) 6.05% (d) 5.08%
23. 20.0 g of a magnesium carbonate sample decomposes on heating to give carbon dioxide and 8.0 g magnesium oxide. What will be the percentage purity of magnesium carbonate in the sample ?
 (a) 75 (b) 96 (c) 60 (d) 84
24. A sample of AlF_3 contains $3.0 \times 10^{24} F^-$ ions. The number of formula unit of this sample are
 (a) 9×10^{24} (b) 3×10^{24}
 (c) 0.75×10^{24} (d) 1.0×10^{24}
25. Read the following and choose the incorrect statements.
 (i) Both weight and mass are same quantities used for measurement of amount of matter present in a substance
 (ii) Mass and weight of a substance vary from one place to another due to change in gravity.
 (iii) SI unit of mass is kilogram and while SI unit of weight is gram.
 (a) (i) and (iii) (b) (ii) and (iii)
 (c) (i) and (ii) (d) All of these
26. Number of atoms in 558.5 grams of Fe (at. wt. of Fe = 55.85 $g\ mol^{-1}$) is
 (a) twice that in 60 g carbon
 (b) 6.023×10^{22}
 (c) half that in 8 g He
 (d) $558.5 \times 6.023 \times 10^{23}$
27. What is the mass of precipitate formed when 50 mL of 16.9% solution of $AgNO_3$ is mixed with 50 mL of 5.8% $NaCl$ solution ?
 (Ag = 107.8, N = 14, O = 16, Na = 23, Cl = 35.5)
 (a) 28 g (b) 3.5 g (c) 7 g (d) 14 g
28. Which of the following option represents correct limiting reagents in reactions (i), (ii) and (iii) respectively.
 (i) $C + O_2 \rightarrow CO_2$
 (26g) (20g)
 (ii) $N_2 + 3H_2 \rightarrow 2NH_3$
 (60g) (80g)
 (iii) $P_4 + 3O_2 \rightarrow P_4O_6$
 (100g) (200g)
 (a) C, N_2 , O_2 (b) C, N_2 , P_4
 (c) O_2 , H_2 , P_4 (d) O_2 , N_2 , P_4
29. A compound made up of two elements A and B is found to contain 25% A (atomic mass = 12.5) and 75% B (atomic mass = 37.5). The simplest formula of the compound is
 (a) AB (b) AB_2 (c) AB_3 (d) A_3B
30. On analysis a certain compound was found to contain iodine and oxygen in the ratio of 254 g of iodine (atomic mass 127) and 80 g oxygen (at mass = 16). What is the formula of the compound.
 (a) IO (b) I_2O
 (c) I_5O_3 (d) I_2O_5
31. The following equation is a completely balanced equation :
 $3Sn + 12HCl + 4HNO_3 \longrightarrow 3SnCl_4 + 4NO + 8H_2O$
 In the above reaction, the number of equivalent per formula weight of HNO_3 is
 (a) 3 (b) 4 (c) 1 (d) 2

**RESPONSE
GRID**

- | | | | | |
|------------------|------------------|------------------|------------------|------------------|
| 20. (a)(b)(c)(d) | 21. (a)(b)(c)(d) | 22. (a)(b)(c)(d) | 23. (a)(b)(c)(d) | 24. (a)(b)(c)(d) |
| 25. (a)(b)(c)(d) | 26. (a)(b)(c)(d) | 27. (a)(b)(c)(d) | 28. (a)(b)(c)(d) | 29. (a)(b)(c)(d) |
| 30. (a)(b)(c)(d) | 31. (a)(b)(c)(d) | | | |

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32. In a compound C, H and N are present in 9 : 1 : 3.5 by weight. If molecular weight of the compound is 108, then the molecular formula of the compound is :
 (a) $C_2H_6N_2$ (b) C_3H_4N (c) $C_6H_8N_2$ (d) $C_9H_{12}N_3$
33. Arrange the numbers in increasing no. of significant figures. 0.002600, 2.6000, 2.6, 0.260
 (a) $2.6 < 0.260 < 0.002600 < 2.6000$
 (b) $2.6000 < 2.6 < 0.002600 < 0.260$
 (c) $0.260 < 2.6 < 0.002600 < 2.6000$
 (d) $0.002600 < 0.260 < 2.6 < 2.6000$
34. How many moles of lead (II) chloride will be formed from a reaction between 6.5 g of PbO and 3.2 g of HCl ?
 (a) 0.044 (b) 0.333
 (c) 0.011 (d) 0.029
35. Equal weights of NaCl and KCl are dissolved separately in equal volumes of solutions. Molarity of the two solutions will be :
 (a) equal
 (b) that of NaCl will be less than that of KCl
 (c) that of NaCl will be more than that of KCl solution
 (d) that of NaCl will be about half of that of KCl solution
36. Gastric juice contains 3.0 g of HCl per litre. If a person produces 2.5 litre of gastric juice per day. How many antacid tablets each containing 400 mg of $Al(OH)_3$ are needed to neutralize all the HCl produced in one day ?
 (a) 18 (b) 14 (c) 20 (d) 17
37. Which of the following is the correct empirical and molecular formulae of a compound, if the molecular mass of a compound is 80 and compound contains 60% of C, 5% of H and 35% of N ?
 (a) C_2H_2N ; $C_4H_4N_2$ (b) $C_3H_4N_2$; $C_6H_8N_4$
 (c) $C_2H_4N_2$; $C_4H_8N_4$ (d) C_2H_2N ; C_2H_2N
38. A gas mixture of 3 litres of propane (C_3H_8) and butane (C_4H_{10}) on complete combustion at $25^\circ C$ produced 10 litre CO_2 . Find out the composition of gas mixture (Propane : Butane)
 (a) 2 : 1 (b) 1 : 2
 (c) 1.5 : 1.5 (d) 0.5 : 2.5
39. Arrange the following in the order of increasing mass (atomic mass: O = 16, Cu = 63, N = 14)
 I. one atom of oxygen
 II. one atom of nitrogen
 III. 1×10^{-10} mole of oxygen
 IV. 1×10^{-10} mole of copper
 (a) $II < I < III < IV$ (b) $I < II < III < IV$
 (c) $III < II < IV < I$ (d) $IV < II < III < I$
40. When 30 litres of H_2 and 30 litres of N_2 are reacted NH_3 is formed and the yield is only 50%. The composition of the gaseous mixture will be
 (a) 5L of N_2 , 5L of H_2 and 5L of NH_3 .
 (b) 5L of N_2 , 10L of H_2 and 10L of NH_3 .
 (c) 10L of N_2 , 15L of H_2 and 5L of NH_3 .
 (d) 5L of N_2 , 15L of H_2 and 10L of NH_3 .
41. How many moles of magnesium phosphate, $Mg_3(PO_4)_2$ will contain 0.25 mole of oxygen atoms?
 (a) 1.25×10^{-2} (b) 2.5×10^{-2}
 (c) 0.02 (d) 3.125×10^{-2}
42. 1.12 ml of a gas is produced at S.T.P. by the action of 4.12 mg of alcohol ROH with methyl magnesium Iodide. The molecular mass of alcohol is
 (a) 16.0 (b) 41.2
 (c) 82.4 (d) 156.0
43. If 224 mL of a triatomic gas has a mass of 1 g at 273K and 1 atmospheric pressure then the mass of one atom is
 (a) 8.30×10^{-23} g (b) 2.08×10^{-23} g
 (c) 5.53×10^{-23} g (d) 6.24×10^{-23} g
44. A compound contains atoms of three elements as A, B and C. If the oxidation number of A is +2, B is +5 and that of C is -2, the possible formula of the compound is
 (a) $A_3(B_4C)_2$ (b) $A_3(BC_4)_2$
 (c) ABC_2 (d) $A_2(BC_3)_2$
45. 5 moles of SO_2 and 5 moles of O_2 react to form SO_3 . Number of moles left in total when only 60% SO_2 is used is
 (a) 6.5 (b) 10
 (c) 8 (d) 8.5

**RESPONSE
GRID**

- | | | | | | |
|--|------------------|------------------|------------------|------------------|------------------|
| | 32. (a)(b)(c)(d) | 33. (a)(b)(c)(d) | 34. (a)(b)(c)(d) | 35. (a)(b)(c)(d) | 36. (a)(b)(c)(d) |
| | 37. (a)(b)(c)(d) | 38. (a)(b)(c)(d) | 39. (a)(b)(c)(d) | 40. (a)(b)(c)(d) | 41. (a)(b)(c)(d) |
| | 42. (a)(b)(c)(d) | 43. (a)(b)(c)(d) | 44. (a)(b)(c)(d) | 45. (a)(b)(c)(d) | |

Space for Rough Work