

Solubility table

The table below provides information on the variation of solubility of different substances (mostly inorganic compounds) in water with temperature, at 1 atmosphere pressure. Units of solubility are given in grams per 100 millilitres of water (g/100 ml), unless shown otherwise. The substances are listed in alphabetical order.

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A

Substance	Formula	0 °C	10 °C	15 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Actinium(III) hydroxide</u>	<u>Ac(OH)₃</u>				0.0021								
<u>Aluminium chloride</u>	<u>AlCl₃</u>	43.9	44.9		45.8	46.6	47.3		48.1		48.6		49.0
<u>Aluminium fluoride</u>	<u>AlF₃</u>	0.57	0.56		0.67	0.78	0.91		1.1		1.32		1.72
<u>Aluminium hydroxide</u>	<u>Al(OH)₃</u>				2.262×10 ⁻⁸								
<u>Aluminium nitrate</u>	<u>Al(NO₃)₃</u>	60	66.7		73.9	81.8	88.7	96.0	106	120	132	153	160
<u>Aluminium perchlorate</u>	<u>Al(ClO₄)₃</u>	122	128		133								
<u>Aluminium sulfate</u>	<u>Al₂(SO₄)₃</u>	31.2	33.5		36.4	40.4	45.8	52.2	59.2	66.2	73	80.8	89.0
<u>Ammonia (ml/ml)</u>	<u>NH₃</u>	1176	900		702	565	428	333	252	188	138	100	88
<u>Ammonium acetate</u>	<u>NH₄C₂H₃O₂</u>	102			143		204		311		533		
<u>Ammonium azide</u>	<u>NH₄N₃</u>	16			25.3		37.1						
<u>Ammonium benzoate</u>	<u>NH₄C₇H₅O₂</u>		19.6		21.3								83
<u>Ammonium bicarbonate</u>	<u>NH₄HCO₃</u>	11.9	16.1		21.7	28.4	36.6		59.2		109	<u>dec</u>	
<u>Ammonium bromide</u>	<u>NH₄Br</u>	60.6	68.1		76.4	83.2	91.2	99.2	108	117	125	135	145
<u>Ammonium carbonate</u>	<u>(NH₄)₂CO₃·H₂O</u>	55.8			100				<u>dec</u>				
<u>Ammonium chlorate</u>	<u>NH₄ClO₃</u>				28.7								
<u>Ammonium chloride</u>	<u>NH₄Cl</u>	29.4	33.2		37.2	41.4	45.8	50.4	55.3	60.2	65.6	71.2	77.3
<u>Ammonium hexachloroplatinate</u>	<u>(NH₄)₂PtCl₆</u>	0.289	0.374		0.499	0.637	0.815		1.44		2.16	2.61	3.36
<u>Ammonium chromate</u>	<u>(NH₄)₂CrO₄</u>	25	29.2		34	39.3	45.3	51.9	59.0	71.2	76.1		
<u>Ammonium dichromate</u>	<u>(NH₄)₂Cr₂O₇</u>	18.2	25.5		35.6	46.5	58.5	71.4	86.0		115		156
<u>Ammonium dihydrogen arsenate</u>	<u>NH₄H₂AsO₄</u>	33.7			48.7		63.8		83		107	122	
<u>Ammonium dihydrogen phosphate</u>	<u>NH₄H₂PO₄</u>	22.7	29.5		37.4	46.4	56.7	69.0	82.5	98.6	118.3	142.8	173.2
<u>Ammonium fluoride</u>	<u>NH₄F</u>	100			85.3 (25 °C)								
<u>Ammonium fluorosilicate</u>	<u>(NH₄)₂SiF₆</u>	12.28	16.41		18.6	25.0	31.6	35.4	40.4	44.9	48.1(75°C)		61.0
<u>Ammonium formate</u>	<u>NH₄HCO₂</u>	102			143		204		311		533		
<u>Ammonium hydrogen phosphate</u>	<u>(NH₄)₂HPO₄</u>	42.9	62.9		68.9	75.1	81.8	89.2	97.2	106	110	112	121
<u>Ammonium hydrogen sulfate</u>	<u>NH₄HSO₄</u>				100								
<u>Ammonium hydrogen tartrate</u>	<u>NH₄HC₄H₄O₆</u>		1.88		2.7								
<u>Ammonium iodate</u>	<u>NH₄IO₃</u>			2.6									14.5
<u>Ammonium iodide</u>	<u>NH₄I</u>	155	163		172	182	191	200	209	219	229		250
<u>Ammonium nitrate</u>	<u>NH₄NO₃</u>	118	150		192	242	297	344	421	499	580	740	871
<u>Ammonium orthoperiodate</u>	<u>(NH₄)₅IO₆</u>				2.7								
<u>Ammonium oxalate</u>	<u>(NH₄)₂C₂O₄</u>	2.2	3.21		4.45	6.09	8.18	10.3	14.0		22.4	27.9	34.7
<u>Ammonium perchlorate</u>	<u>NH₄ClO₄</u>	11.56	16.4		20.85		30.58		39.05		48.19		57.01
<u>Ammonium permanganate</u>	<u>NH₄MnO₄</u>			8.0					<u>dec</u>				
<u>Ammonium perrhenate</u>	<u>NH₄ReO₄</u>	2.8			6.2		12.0		20.7		32.3	39.1	
<u>Ammonium phosphate</u>	<u>(NH₄)₃PO₄</u>	9.40			20.3			37.7					
<u>Ammonium</u>	<u>(NH₄)₂SeO₄</u>	96	105		115	126	143		192				

<u>selenate</u>													
<u>Ammonium sulfate</u>	$(\text{NH}_4)_2\text{SO}_4$	70.6	73		75.4	78.1	81.2	84.3	87.4		94.1		103
<u>Ammonium aluminium sulfate</u>	$\text{NH}_4\text{Al}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$	2.4	5.0		7.4	10.5	14.6	19.6	26.7	37.7	53.9	98.2	121
<u>Ammonium sulfite</u>	$(\text{NH}_4)_2\text{SO}_3$	47.9	54		60.8	68.8	78.4		104		144	150	153
<u>Ammonium tartrate</u>	$(\text{NH}_4)_2\text{C}_4\text{H}_4\text{O}_6$	45	55		63	70.5	76.5		86.9				
<u>Ammonium thiocyanate</u>	NH_4SCN	120	144		170	208	234	235	346				
<u>Ammonium thiosulfate</u>	$(\text{NH}_4)_2\text{S}_2\text{O}_3$				173		205				269		
<u>Ammonium vanadate</u>	NH_4VO_3				0.48	0.84	1.32	1.78	2.42	3.05			7.0
<u>Aniline</u>	$\text{C}_6\text{H}_7\text{N}$				3.6								
<u>Antimony trifluoride</u>	SbF_3	385			444	562	<u>dec</u>						
<u>Antimony sulfide</u>	Sb_2S_3				1.8×10^{-4}								
<u>Antimony trichloride</u>	SbCl_3	602			910	1090	1370	1917	4531	<u>dec</u>			
<u>Argon (Unit:ml/ml)</u>	Ar	0.056	0.0405		0.0336	0.0288	0.0252	0.0223					
<u>Arsenic pentasulfide</u>	As_2S_5	0.0014											
<u>Arsenic pentoxide</u>	As_2O_5	59.5	62.1		65.8	70.6	71.2		73.0		75.1		76.7
<u>Arsenious sulfide</u>	As_2S_3				0.0004								
<u>Arsenic trioxide</u>	As_2O_3	1.21	1.58		1.80		2.93	3.43	4.44	5.37	5.89	6.55	9
<u>Arsine (Unit:ml/ml)</u>	AsH_3				0.2								

B

Calcium carbonate (Calcite)	<u>CaCO₃-Calcite</u>			6.17×10 ⁻⁴							
Calcium chlorate	<u>Ca(ClO₃)₂</u>			209							
Calcium chloride	<u>CaCl₂</u>	59.5	64.7	74.5	100	128		137		147	154
Calcium chromate	<u>CaCrO₄</u>	4.5		2.25	1.83	1.49		0.83			
Calcium citrate	<u>Ca₃(C₆H₅O₇)₂</u>			0.095 (25 °C)							
Monocalcium phosphate	<u>Ca(H₂PO₄)₂</u>			1.8							
Calcium fluoride	<u>CaF₂</u>	0.008575									
Calcium fluorosilicate	<u>CaSiF₆</u>			0.518							
Calcium formate	<u>Ca(HCO₂)₂</u>	16.1		16.6		17.1		17.5		17.9	18.4
Dicalcium phosphate	<u>CaHPO₄</u>			0.004303							
Calcium hydroxide	<u>Ca(OH)₂</u>	0.189	0.182	0.173	0.16	0.141		0.121		0.086	0.076
Calcium iodate	<u>Ca(IO₃)₂</u>	0.09		0.24	0.38	0.52		0.65		0.66	0.67
Calcium iodide	<u>CaI₂</u>	64.6		66	67.6	70.8		74		78	81
Calcium molybdate	<u>CaMoO₄</u>			0.004099							
Calcium nitrate	<u>Ca(NO₃)₂</u>			121.2							
Calcium nitrate tetrahydrate	<u>Ca(NO₃)₂·4H₂O</u>	102	115	129	152	191				358	363
Calcium nitrite	<u>Ca(NO₂)₂·4H₂O</u>	63.9		84.5	104			134		151	166
Calcium oxalate	<u>CaC₂O₄</u>			6.7×10 ⁻⁴							0.0014
Calcium oxide	<u>CaO</u>										5.7
Calcium perchlorate	<u>Ca(ClO₄)₂</u>			188							
Calcium permanganate	<u>Ca(MnO₄)₂</u>			338							
Calcium phosphate	<u>Ca₃(PO₄)₂</u>			0.002							
Calcium selenate	<u>CaSeO₄·2H₂O</u>	9.73	9.77	9.22	8.79	7.14					
Calcium sulfate	<u>CaSO₄·2H₂O</u>	0.223	0.244	0.255	0.264	0.265		0.244		0.234	0.205
Calcium tungstate	<u>CaWO₄</u>			0.002387							
Carbon dioxide	<u>CO₂</u>			0.1782							
Carbon monoxide	<u>CO</u>			0.0026							
Cerium(III) acetate	<u>Ce(C₂H₃O₂)₃</u>			26							
Cerium(III) chloride	<u>CeCl₃</u>			100							
Cerium(III) hydroxide	<u>Ce(OH)₃</u>			9.43×10 ⁻⁵							
Cerium(III) iodate	<u>Ce(IO₃)₃</u>			0.123							
Cerium(III) nitrate	<u>Ce(NO₃)₃</u>			234							
Cerium(III) phosphate	<u>CePO₄</u>			7.434×10 ⁻¹¹							
Cerium(III) selenate	<u>Ce₂(SeO₄)₃</u>	39.5	37.2	35.2	33.2	32.6		13.7		4.6	
Cerium(III) sulfate	<u>Ce₂(SO₄)₃·2H₂O</u>	21.4		9.84	7.24	5.63		3.87			
Cerium(IV) hydroxide	<u>Ce(OH)₄</u>			1.981×10 ⁻⁵							
Chromium(III) nitrate	<u>Cr(NO₃)₃</u>	108	124	130	152						
Chromium(III) perchlorate	<u>Cr(ClO₄)₃</u>	104	123	130							
Chromium(III) sulfate	<u>Cr₂(SO₄)₃·18H₂O</u>			220							
Chromium(VI) oxide	<u>CrO₃</u>	61.7		63							67
Cobalt(II) bromate	<u>Co(BrO₃)₂·6H₂O</u>			45.5							
Cobalt(II) bromide	<u>CoBr₂</u>	91.9		112	128	163		227		241	257
Cobalt(II) chlorate	<u>Co(ClO₃)₂</u>	135	162	180	195	214		316			
Cobalt(II) chloride	<u>CoCl₂</u>	43.5	47.7	52.9	59.7	69.5		93.8		97.6	101
Cobalt(II) fluoride	<u>CoF₂</u>			1.36							
Cobalt(II) fluorosilicate	<u>CoSiF₆·6H₂O</u>			118							
Cobalt(II) iodate	<u>Co(IO₃)₂·2H₂O</u>			1.02	0.9	0.88		0.82		0.73	0.7

Cobalt(II) iodide	<u>CoI₂</u>			203							
Cobalt(II) nitrate	<u>Co(NO₃)₂</u>	84	89.6	97.4	111	125		174		204	300
Cobalt(II) nitrite	<u>Co(NO₂)₂</u>	0.076	0.24	0.4	0.61	0.85					
Cobalt oxalate	<u>CoC₂O₄·2H₂O</u>			2.6972×10 ⁻⁹							
Cobalt(II) perchlorate	<u>Co(ClO₄)₂</u>			104							
Cobalt(II) sulfate	<u>CoSO₄</u>	25.5	30.5	36.1	42	48.8		55		53.8	45.3 38.9
Copper(I) chloride	<u>CuCl</u>			0.0099							
Copper(I) cyanide	<u>CuCN</u>			1.602×10 ⁻⁹							
Copper(I) hydroxide	<u>CuOH</u>			8.055×10 ⁻⁷							
Copper(I) iodide	<u>CuI</u>			0.0042							
Copper(I) sulfide	<u>Cu₂S</u>			1.361×10 ⁻¹⁵							
Copper(I) thiocyanate	<u>CuSCN</u>			8.427×10 ⁻⁷							
Copper(II) bromide	<u>CuBr₂</u>	107	116	126	128	131					
Copper(II) carbonate	<u>CuCO₃</u>			1.462×10 ⁻⁴							
Copper(II) chlorate	<u>Cu(ClO₃)₂</u>			242							
Copper(II) chloride	<u>CuCl₂</u>	68.6	70.9	73	77.3	87.6		96.5		104	108 120
Copper(II) chromate	<u>CuCrO₄</u>			0.03407							
Copper(II) fluoride	<u>CuF₂</u>			0.075							
Copper(II) fluorosilicate	<u>CuSiF₆</u>	73.5	76.5	81.6	84.1	91.2				93.2	
Copper(II) formate	<u>Cu(HCO₂)₂</u>			12.5							
Copper(II) hydroxide	<u>Cu(OH)₂</u>			1.722×10 ⁻⁶							
Copper(II) iodate	<u>Cu(IO₃)₂·2H₂O</u>			0.109							
Copper(II) nitrate	<u>Cu(NO₃)₂</u>	83.5	100	125	156	163		182		208	222 247
Copper oxalate	<u>CuC₂O₄·2H₂O</u>			2.1627×10 ⁻¹⁰							
Copper(II) perchlorate	<u>Cu(ClO₄)₂</u>				146						
Copper(II) selenate	<u>CuSeO₄</u>	12	14.5	17.5	21	25.2		36.5		53.7	
Copper(II) selenite	<u>CuSeO₃</u>			0.002761							
Copper(II) sulfate	<u>CuSO₄·5H₂O</u>	23.1	27.5	32	37.8	44.6		61.8		83.8	114
Copper(II) sulfide	<u>CuS</u>			2.41×10 ⁻¹⁷							

D

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Dysprosium(III) chromate	<u>Dy₂(CrO₄)₃·10H₂O</u>			0.663								
Dysprosium(III) sulfate	<u>Dy₂(SO₄)₃·8H₂O</u>			4.83								

E

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Erbium(III) hydroxide	<u>Er(OH)₃</u>			1.363×10 ⁻⁵								
Erbium(III) sulfate	<u>Er₂(SO₄)₃</u>			13.79								
Erbium(III) sulfate octahydrate	<u>Er₂(SO₄)₃·8H₂O</u>			16.00		6.53						
Europium(III) hydroxide	<u>Eu(OH)₃</u>			1.538×10 ⁻⁵								
Europium(III) sulfate	<u>Eu₂(SO₄)₃·8H₂O</u>			2.56								

F-G

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Ferrous ammonium sulfate</u>	<u>(NH₄)₂Fe(SO₄)₂·6H₂O</u>			26.9						73		
<u>Fructose</u>	<u>C₆H₁₂O₆</u>			375.0		538.0						
<u>Gadolinium(III) acetate</u>	<u>Gd(C₂H₃O₂)₃·4H₂O</u>			11.6								
<u>Gadolinium(III) bicarbonate</u>	<u>Gd(HCO₃)₃</u>			5.61								
<u>Gadolinium(III) bromate</u>	<u>Gd(BrO₃)₃·9H₂O</u>	50.2	70.1	95.6	126	166						
<u>Gadolinium(III) hydroxide</u>	<u>Gd(OH)₃</u>			1.882×10 ⁻⁵								
<u>Gadolinium(III) sulfate</u>	<u>Gd₂(SO₄)₃</u>	3.98	3.3	2.6	2.32							
<u>D-Galactose</u>	<u>C₆H₁₂O₆</u>			10.3								68.3
<u>Gallium chloride</u>	<u>GaCl₃</u>			180								
<u>Gallium hydroxide</u>	<u>Ga(OH)₃</u>			8.616×10 ⁻⁹								
<u>Gallium oxalate</u>	<u>Ga₂(C₂O₄)₃·4H₂O</u>			0.4								
<u>Gallium selenate</u>	<u>Ga₂(SeO₄)₃·16H₂O</u>			18.1								
<u>D-Glucose</u>	<u>C₆H₁₂O₆</u>			90								
<u>Gold(III) chloride</u>	<u>AuCl₃</u>			68								
<u>Gold(V) oxalate</u>	<u>Au₂(C₂O₄)₅</u>			0.258								

H

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Hafnium(III) hydroxide</u>	<u>Hf(OH)₃</u>			4.503×10 ⁻⁴								
<u>Hafnium(IV) hydroxide</u>	<u>Hf(OH)₄</u>			4.503×10 ⁻⁶								
<u>Helium</u>	<u>He</u>			0.6								
<u>Holmium(III) hydroxide</u>	<u>Ho(OH)₃</u>			2.519×10 ⁻⁵								
<u>Holmium(III) sulfate</u>	<u>Ho₂(SO₄)₃·8H₂O</u>			8.18	6.1	4.52						
<u>Hydrogen chloride</u>	<u>HCl</u>	81	75	70	65.5	61	57.5	53	50	47	43	40
<u>Hydrogen sulfide</u>	<u>H₂S</u>			0.33								

I

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Indium(III) bromide</u>	<u>InBr₃</u>			571								
<u>Indium(III) chloride</u>	<u>InCl₃</u>		210	212								
<u>Indium(III) fluoride</u>	<u>InF₃</u>			11.2								
<u>Indium(III) hydroxide</u>	<u>In(OH)₃</u>			3.645×10 ⁻⁸								
<u>Indium(III) iodate</u>	<u>In(IO₃)₃</u>			0.067								
<u>Indium(III) sulfide</u>	<u>In₂S₃</u>			2.867×10 ⁻¹⁴								
<u>Iron(II) bromide</u>	<u>FeBr₂</u>	101	109	117	124	133		144		168	176	184
<u>Iron(II) carbonate</u>	<u>FeCO₃</u>			6.554×10 ⁻⁵								
<u>Iron(II) chloride</u>	<u>FeCl₂</u>	49.7	59	62.5	66.7	70		78.3		88.7	92.3	94.9
<u>Iron(II) fluorosilicate</u>	<u>FeSiF₆.6H₂O</u>	72.1	74.4		77			84		88		100
<u>Iron(II) hydroxide</u>	<u>Fe(OH)₂</u>			5.255×10 ⁻⁵								
<u>Iron(II) nitrate</u>	<u>Fe(NO₃)₂.6H₂O</u>	113	134									
<u>Iron(II) oxalate</u>	<u>FeC₂O₄.2H₂O</u>			0.008								
<u>Iron(II) perchlorate</u>	<u>Fe(ClO₄)₂.6H₂O</u>			299								
<u>Iron(II) sulfate</u>	<u>FeSO₄</u>			28.8		40	48	60	73.3		101	79.9
<u>Iron(III) arsenate</u>	<u>FeAsO₄</u>			1.47×10 ⁻⁹								
<u>Iron(III) chloride</u>	<u>FeCl₃.6H₂O</u>	74.4		91.8	107							
<u>Iron(III) fluoride</u>	<u>FeF₃</u>			0.091								
<u>Iron(III) hydroxide</u>	<u>Fe(OH)₃</u>			2.097×10 ⁻⁹								
<u>Iron(III) iodate</u>	<u>Fe(IO₃)₃</u>			0.36								
<u>Iron(III) nitrate</u>	<u>Fe(NO₃)₃.9H₂O</u>	112		138		175						
<u>Iron(III) perchlorate</u>	<u>Fe(ClO₄)₃</u>	289		368	422	478		772				
<u>Iron(III) sulfate</u>	<u>Fe₂(SO₄)₃.9H₂O</u>											

L

<u>phosphate</u>												
<u>Lithium fluoride</u>	<u>LiF</u>			0.27	0.135							
<u>Lithium fluorosilicate</u>	<u>Li₂SiF₆·2H₂O</u>			73								
<u>Lithium formate</u>	<u>LiHCO₂</u>	32.3	35.7	39.3	44.1	49.5		64.7		92.7	116	138
<u>Lithium hydrogen phosphite</u>	<u>Li₂HPO₃</u>	4.43			9.97	7.61		7.11				6.03
<u>Lithium hydroxide</u>	<u>LiOH</u>	12.7	12.7	12.8	12.9	13.0	13.3	13.8		15.3		17.5
<u>Lithium iodide</u>	<u>LiI</u>	151	157	165	171	179		202		435	440	481
<u>Lithium molybdate</u>	<u>Li₂MoO₄</u>	82.6		79.5	79.5	78						73.9
<u>Lithium nitrate</u>	<u>LiNO₃</u>	53.4	60.8	70.1	138	152		175				
<u>Lithium nitrite</u>	<u>LiNO₂</u>	70.9	82.5	96.8	114	133		177		233	272	324
<u>Lithium oxalate</u>	<u>Li₂C₂O₄</u>			8								
<u>Lithium perchlorate</u>	<u>LiClO₄</u>	42.7	49	56.1	63.6	72.3		92.3		128	151	
<u>Lithium permanganate</u>	<u>LiMnO₄</u>			71.4								
<u>Lithium phosphate</u>	<u>Li₃PO₄</u>			0.039								
<u>Lithium selenide</u>	<u>Li₂Se</u>			57.7								
<u>Lithium selenite</u>	<u>Li₂SeO₃</u>	25	23.3	21.5	19.6	17.9		14.7		11.9	11.1	9.9
<u>Lithium sulfate</u>	<u>Li₂SO₄</u>	36.1	35.5	34.8	34.2	33.7		32.6		31.4	30.9	
<u>Lithium tartrate</u>	<u>Li₂C₄H₄O₆</u>	42	31.8	27.1	26.6	27.2		29.5				
<u>Lithium thiocyanate</u>	<u>LiSCN</u>			114	131	153						
<u>Lithium vanadate</u>	<u>LiVO₃</u>	2.5		4.82	6.28	4.38		2.67				
<u>Lutetium(III) hydroxide</u>	<u>Lu(OH)₃</u>			1.164×10 ⁻⁵								
<u>Lutetium(III) sulfate</u>	<u>Lu₂(SO₄)₃·8H₂O</u>			57.9								

M

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Magnesium acetate	$Mg(C_2H_3O_2)_2$	56.7	59.7	53.4	68.6	75.7		118				
Magnesium benzoate	$Mg(C_7H_5O_2)_2 \cdot H_2O$					5						
Magnesium bromate	$Mg(BrO_3)_2 \cdot 6H_2O$					58						
Magnesium bromide	$MgBr_2$	98	99	101	104	106		112				125
Magnesium carbonate	$MgCO_3$			0.039								
Magnesium chlorate	$Mg(ClO_3)_2$	114	123	135	155	178		242			268	
Magnesium chloride	$MgCl_2$	52.9	53.6	54.6	55.8	57.5		61		66.1	69.5	73.3
Magnesium chromate	$MgCrO_4 \cdot 7H_2O$			137								
Magnesium fluoride	MgF_2			0.007325								
Magnesium fluorosilicate	$MgSiF_6$	26.3		30.8		34.9		44.4				
Magnesium formate	$Mg(HCO_2)_2$	14	14.2	14.4	14.9	15.9		17.9		20.5	22.2	22.9
Magnesium hydroxide	$Mg(OH)_2$			9.628×10^{-4}								0.004
Magnesium iodate	$Mg(IO_3)_2$		7.2	8.6	10	11.7		15.2		15.5	15.6	
Magnesium iodide	MgI_2	120		140		173				186		
Magnesium molybdate	$MgMoO_4$			13.7								
Magnesium nitrate	$Mg(NO_3)_2$	62.1	66	69.5	73.6	78.9		78.9		91.6	106	
Magnesium oxalate	MgC_2O_4			0.104								
Magnesium perchlorate	$Mg(ClO_4)_2$			49.6								
Magnesium phosphate	$Mg_3(PO_4)_2$			2.588×10^{-4}								
Magnesium selenate	$MgSeO_4$	20	30.4	38.3	44.3	48.6		55.8				
Magnesium selenite	$MgSeO_3$			0.05454								
Magnesium sulfate	$MgSO_4$	25.5	30.4	35.1	39.7	44.7	50.4	54.8	59.2	54.8	52.9	50.2
Magnesium sulfite	$MgSO_3 \cdot 6H_2O$			$0.52(25\text{ °C})^{[1]}$								
Magnesium thiosulfate	MgS_2O_3			50								
Maltose	$C_{12}H_{22}O_{11}$			108								
D-Mannose	$C_6H_{12}O_6$			248								
Manganese(II) bromide	$MnBr_2$	127	136	147	157	169		197		225	226	228
Manganese(II) carbonate	$MnCO_3$			4.877×10^{-5}								
Manganese(II) chloride	$MnCl_2$	63.4	68.1	73.9	80.8	88.5		109		113	114	115
Manganese(II) ferrocyanide	$Mn_2Fe(CN)_6$			0.001882								
Manganese(II) fluoride	MnF_2			0.96		0.67		0.44				0.48
Manganese(II) fluorosilicate	$MnSiF_6 \cdot 6H_2O$			140								
Manganese(II) hydroxide	$Mn(OH)_2$			3.221×10^{-4}								
Manganese(II) nitrate	$Mn(NO_3)_2$	102	118	139	206							
Manganese(II) oxalate	$MnC_2O_4 \cdot 2H_2O$	0.02	0.024	0.028	0.033							
Manganese(II) sulfate	$MnSO_4$	52.9	59.7	62.9	62.9	60		53.6		45.6	40.9	35.3
Mercury(I) azide	$Hg_2(N_3)_2$			0.02727								
Mercury(I) bromide	Hg_2Br_2			1.352×10^{-6}								
Mercury(I) carbonate	Hg_2CO_3			4.351×10^{-7}								
Mercury(I) chloride	Hg_2Cl_2			3.246×10^{-5}								
Mercury(I) chromate	Hg_2CrO_4			0.002313								
Mercury(I) cyanide	$Hg_2(CN)_2$			2.266×10^{-12}								
Mercury(I) perchlorate	$Hg_2(ClO_4)_2$	282	325	407	455		499		541		580	
Mercury(I) sulfate	Hg_2SO_4			0.04277								
Mercury(II) acetate	$Hg(C_2H_3O_2)_2$			25								
Mercury(II) benzoate	$Hg(C_7H_5O_2)_2 \cdot H_2O$			1.1								
Mercury(II) bromate	$Hg(BrO_3)_2 \cdot 2H_2O$			0.08								
Mercury(II) bromide	$HgBr_2$	0.3	0.4	0.56	0.66	0.91		1.68		2.77		4.9

Mercury(II) chlorate	$\text{Hg}(\text{ClO}_3)_2$			25								
Mercury(II) chloride	HgCl_2	3.63	4.82	6.57	8.34	10.2		16.3		30		61.3
Mercury(II) cyanide	$\text{Hg}(\text{CN})_2$			9.3								
Mercury(II) iodate	$\text{Hg}(\text{IO}_3)_2$			0.002372								
Mercury(II) iodide	HgI_2			0.006								
Mercury(II) oxalate	HgC_2O_4			0.011								
Mercury(II) sulfide	HgS			2.943×10^{-25}								
Mercury(II) thiocyanate	$\text{Hg}(\text{SCN})_2$			0.063								

N

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Neodymium(III) acetate	$\text{Nd}(\text{C}_2\text{H}_3\text{O}_2)_3 \cdot \text{H}_2\text{O}$			26.2								
Neodymium(III) bromate	$\text{Nd}(\text{BrO}_3)_3$	43.9	59.2	75.6	95.2	116						
Neodymium(III) chloride	NdCl_3		96.7	98	99.6	102		105				
Neodymium(III) molybdate	$\text{Nd}_2(\text{MoO}_4)_3$				0.0019							
Neodymium(III) nitrate	$\text{Nd}(\text{NO}_3)_3$	127	133	142	145	159		211				
Neodymium(III) selenate	$\text{Nd}_2(\text{SeO}_4)_3$	45.2	44.6	41.8	39.9	39.9		43.9		7	3.3	
Neodymium(III) sulfate	$\text{Nd}_2(\text{SO}_4)_3$	13	9.7	7.1	5.3	4.1		2.8		2.2	1.2	
Nickel(II) acetate	$\text{Ni}(\text{C}_2\text{H}_3\text{O}_2)_2$											
Nickel(II) bromate	$\text{Ni}(\text{BrO}_3)_2 \cdot 6\text{H}_2\text{O}$			28								
Nickel(II) bromide	NiBr_2	113	122	131	138	144		153		154		155
Nickel(II) carbonate	NiCO_3			9.643×10^{-4}								
Nickel(II) chlorate	$\text{Ni}(\text{ClO}_3)_2$	111	120	133	155	181		221		308		
Nickel(II) chloride	NiCl_2	53.4	56.3	66.8	70.6	73.2		81.2		86.6		87.6
Nickel(II) fluoride	NiF_2		2.55	2.56				2.56			2.59	
Nickel(II) formate	$\text{Ni}(\text{HCO}_2)_2 \cdot 2 \text{H}_2\text{O}$		3.15	3.25								
Nickel(II) hydroxide	$\text{Ni}(\text{OH})_2$			0.013								
Nickel(II) iodate	$\text{Ni}(\text{IO}_3)_2$	0.74		0.062	1.43							
Nickel(II) iodide	NiI_2	124	135	148	161	174		184		187	188	
Nickel(II) nitrate	$\text{Ni}(\text{NO}_3)_2$	79.2		94.2	105	119		158		187	188	
Nickel oxalate	$\text{NiC}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$			0.00118								
Nickel(II) perchlorate	$\text{Ni}(\text{ClO}_4)_2$	105	107	110	113	117						
Nickel(II) pyrophosphate	$\text{Ni}_2\text{P}_2\text{O}_7$			0.001017								
Nickel(II) sulfate	$\text{NiSO}_4 \cdot 6\text{H}_2\text{O}$			44.4	46.6	49.2		55.6		64.5	70.1	76.7
Nitric oxide	NO			0.0056								
Nitrous oxide	N_2O			0.112								

O

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Oxygen at a partial pressure of 21 kPa	O_2	0.00146	0.00113	0.00091	0.00076	0.00065						
Oxalic acid	$\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$	4.96	8.51	13.3	19.9	30.1		62.1		118	168	

P

Potassium thiocyanate	<u>KSCN</u>	177	198	224	255	289		372		492	571	675
Potassium thiosulfate	<u>K₂S₂O₃</u>	96		155	175	205		238		293	312	
Potassium tungstate	<u>K₂WO₄</u>			51.5								
Praseodymium(III) acetate	<u>Pr(C₂H₃O₂)₃·H₂O</u>			32								
Praseodymium(III) bromate	<u>Pr(BrO₃)₃</u>	55.9	73	91.8	114	144						
Praseodymium(III) chloride	<u>PrCl₃</u>			104								
Praseodymium(III) molybdate	<u>Pr₂(MoO₄)₃</u>			0.0015								
Praseodymium(III) nitrate	<u>Pr(NO₃)₃</u>			112	162	178						
Praseodymium(III) sulfate	<u>Pr₂(SO₄)₃</u>	19.8	15.6	12.6	9.89	2.56		5.04		3.5	1.1	0.91

R

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Radium chloride	<u>RaCl₂</u>			19.6								
Radium iodate	<u>Ra(IO₃)₂</u>			0.04								
Radium nitrate	<u>Ra(NO₃)₂</u>			12								
Radium sulfate	<u>RaSO₄</u>			2.1×10 ⁻⁴								
Raffinose	<u>C₁₈H₃₂O₁₆·5H₂O</u>			14								
Rubidium acetate	<u>RbC₂H₃O₂</u>					86						
Rubidium bromate	<u>RbBrO₃</u>				3.6	5.1						
Rubidium bromide	<u>RbBr</u>	90	99	108	119	132		158				
Rubidium chlorate	<u>RbClO₃</u>	2.1	3.1	5.4	8	11.6		22		38	49	63
Rubidium chloride	<u>RbCl</u>	77	84	91	98	104		115		127	133	143
Rubidium chromate	<u>Rb₂CrO₄</u>	62	67.5	73.6	78.9	85.6		95.7				
Rubidium dichromate	<u>Rb₂Cr₂O₇</u>			5.9	10	15.2		32.3				
Rubidium fluoride	<u>RbF</u>			130.6 (18 °C)								
Rubidium fluorosilicate	<u>Rb₂SiF₆</u>			0.157								
Rubidium formate	<u>RbHCO₂</u>		443	554	614	694		900				
Rubidium hydrogen carbonate	<u>RbHCO₃</u>			110								
Rubidium hydroxide	<u>RbOH</u>			180								
Rubidium iodate	<u>RbIO₃</u>			1.96								
Rubidium iodide	<u>RbI</u>			144								
Rubidium nitrate	<u>RbNO₃</u>	19.5	33	52.9	81.2	117		200		310	374	452
Rubidium perchlorate	<u>RbClO₄</u>	1.09	1.19	1.55	2.2	3.26		6.27		11	15.5	22
Rubidium periodate	<u>RbIO₄</u>			0.648								
Rubidium phosphate	<u>Rb₃PO₄</u>			220								
Rubidium selenate	<u>Rb₂SeO₄</u>			159								
Rubidium sulfate	<u>Rb₂SO₄</u>	37.5	42.6	48.1	53.6	58.5		67.5		75.1	78.6	81.8

S

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
Samarium acetate	$\text{Sm}(\text{C}_2\text{H}_3\text{O}_2)_3 \cdot 3\text{H}_2\text{O}$			15								
Samarium bromate	$\text{Sm}(\text{BrO}_3)_3$	34.2	47.6	62.5	79	98						
Samarium chloride	SmCl_3		92.4	93.4	94.6	96.9						
Samarium sulfate	$\text{Sm}_2(\text{SO}_4)_3 \cdot 8\text{H}_2\text{O}$			2.7	3.1							
Scandium oxalate	$\text{Sc}_2(\text{C}_2\text{O}_4)_3 \cdot 6\text{H}_2\text{O}$			0.006								
Scandium sulfate	$\text{Sc}_2(\text{SO}_4)_3 \cdot 5\text{H}_2\text{O}$			54.6								
Silicon dioxide	SiO_2			0.012								
Silver acetate	$\text{AgC}_2\text{H}_3\text{O}_2$	0.73	0.89	1.05	1.23	1.43		1.93		2.59		
Silver azide	AgN_3			7.931×10^{-4}								
Silver bromate	AgBrO_3		0.11	0.16	0.23	0.32		0.57		0.94	1.33	
Silver bromide	AgBr			1.328×10^{-5}								
Silver carbonate	Ag_2CO_3			0.003489								
Silver chlorate	AgClO_3		10.4	15.3	20.9	26.8						
Silver chloride	AgCl			1.923×10^{-4}			5.2×10^{-5}					
Silver chlorite	AgClO_2			0.248								
Silver chromate	Ag_2CrO_4			0.002157								
Silver cyanide	AgCN			1.467×10^{-7}								
Silver dichromate	$\text{Ag}_2\text{Cr}_2\text{O}_7$			0.159								
Silver fluoride	AgF	85.9	120	172	190	203						
Silver nitrate	AgNO_3	122	167	216	265	311		440		585	652	733
Silver oxalate	$\text{Ag}_2\text{C}_2\text{O}_4$			0.00327								
Silver oxide	Ag_2O			0.0012								
Silver perchlorate	AgClO_4	455	484	525	594	635						793
Silver permanganate	AgMnO_4			0.9								
Silver sulfate	Ag_2SO_4	0.57	0.7	0.8	0.89	0.98		1.15		1.3	1.36	1.41
Silver vanadate	AgVO_3			0.01462								
Sodium acetate	$\text{NaC}_2\text{H}_3\text{O}_2$	36.2	40.8	46.4	54.6	65.6		139		153	161	170
Sodium azide	NaN_3	38.9	39.9	40.8								
Sodium benzoate	$\text{NaC}_7\text{H}_5\text{O}_2$			66								
Sodium borohydride	NaBH_4	25		55		88.5						
Sodium bromate	NaBrO_3	24.2	30.3	36.4	42.6	48.8		62.6		75.7		90.8
Sodium bromide	NaBr	80.2	85.2	90.8	98.4	107		118		120	121	121
Sodium carbonate	Na_2CO_3	7	12.5	21.5	39.7	49		46		43.9	43.9	45.5
Sodium chlorate	NaClO_3	79.6	87.6	95.9	105	115		137		167	184	204
Sodium chloride	NaCl	35.65	35.72	35.89	36.09	36.37	36.69	37.04	37.46	37.93	38.47	38.99
Sodium chromate	Na_2CrO_4	31.7	50.1	84	88	96		115		125		126
Sodium cyanide	NaCN	40.8	48.1	58.7	71.2	dec						
Sodium dichromate	$\text{Na}_2\text{Cr}_2\text{O}_7$	163	172	183	198	215		269		376	405	415
Monosodium phosphate	NaH_2PO_4	56.5	69.8	86.9	107	133		172		211	234	
Sodium fluoride	NaF	3.66		4.06	4.22	4.4		4.68		4.89		5.08
Sodium formate	HCOONa	43.9	62.5	81.2	102	108		122		138	147	160
Sodium hydrogen carbonate	NaHCO_3	7	8.1	9.6	11.1	12.7		16				
Sodium hydroxide	NaOH		98	109	119	129		174				
Sodium iodate	NaIO_3	2.48	4.59	8.08	10.7	13.3		19.8		26.6	29.5	33
Sodium iodide	NaI	159	167	178	191	205		257		295		302
Sodium metabisulfite	$\text{Na}_2\text{S}_2\text{O}_5$	45.1		65.3						88.7		96.3
Sodium metaborate	NaBO_2	16.4	20.8	25.4	31.4	40.4		63.9		84.5		125.2
Sodium molybdate	Na_2MoO_4	44.1	64.7	65.3	66.9	68.6		71.8				
Sodium nitrate	NaNO_3	73	80.8	87.6	94.9	102		122		148		180
Sodium nitrite	NaNO_2	71.2	75.1	80.8	87.6	94.9		111		133		160

Sodium oxalate	$\text{Na}_2\text{C}_2\text{O}_4$	2.69	3.05	3.41	3.81	4.18		4.93		5.71		6.5
Sodium perchlorate	NaClO_4	167	183	201	222	245		288		306		329
Sodium periodate	NaIO_4	1.83	5.6	10.3	19.9	30.4						
Sodium permanganate	NaMnO_4			90								
Sodium phosphate	Na_3PO_4	4.5	8.2	12.1	16.3	20.2		20.9		60	68.1	77
Sodium pyrophosphate	$\text{Na}_4\text{P}_2\text{O}_7$	2.26										
Sodium selenate	Na_2SeO_4	13.3	25.2	26.9	77	81.8		78.6		74.8	73	72.7
Sodium sulfate	Na_2SO_4	4.9	9.1	19.5	40.8	48.8		45.3		43.7	42.7	42.5
Sodium sulfite	Na_2SO_3			27.0								
Sodium tetraborate decahydrate	$\text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$	2	2.3	2.5	4	6	10	15				
Sodium tetraborate pentahydrate	$\text{Na}_2\text{B}_4\text{O}_7 \cdot 5\text{H}_2\text{O}$								20	23	28	35
Sodium tetraborate tetrahydrate	$\text{Na}_2\text{B}_4\text{O}_7 \cdot 4\text{H}_2\text{O}$								17	20	23	28
Sodium tetrphenylborate	$\text{NaB}(\text{C}_6\text{H}_5)_4$			47								
Sodium thiosulfate	$\text{Na}_2\text{S}_2\text{O}_3$	71.5		73		77.6				90.8		97.2
Strontium acetate	$\text{Sr}(\text{C}_2\text{H}_3\text{O}_2)_2$	37	42.9	41.1	39.5	38.3		36.8		36.1	36.2	36.4
Strontium bromate	$\text{Sr}(\text{BrO}_3)_2 \cdot \text{H}_2\text{O}$			30.9								41
Strontium bromide	SrBr_2	85.2	93.4	102	112	123		150		182		223
Strontium carbonate	SrCO_3			0.0011								0.065
Strontium chlorate	$\text{Sr}(\text{ClO}_3)_2$			175								
Strontium chloride	SrCl_2	43.5	47.7	52.9	58.7	65.3		81.8		90.5		101
Strontium chromate	SrCrO_4			0.085	0.090							
Strontium fluoride	SrF_2			1.2×10^{-4}								
Strontium formate	$\text{Sr}(\text{HCO}_2)_2$	9.1	10.6	12.7	15.2	17.8		25		31.9	32.9	34.4
Strontium hydroxide	$\text{Sr}(\text{OH})_2 \cdot 8\text{H}_2\text{O}$	0.91	1.25	1.77	2.64	3.95		8.42		20.2	44.5	91.2
Strontium iodate	$\text{Sr}(\text{IO}_3)_2$			0.19								0.35
Strontium iodide	SrI_2	165		178		192		218		270	365	383
Strontium molybdate	SrMoO_4			0.01107								
Strontium nitrate	$\text{Sr}(\text{NO}_3)_2$	39.5	54.9	70.8	87.6	91.3	92.6	94.0	97.2	99.0	101.1	
Strontium perchlorate ^[2]	$\text{Sr}(\text{ClO}_4)_2$	233.8	258.7	291.7	327.5	363.9						
Strontium selenate	SrSeO_4			0.656								
Strontium sulfate	SrSO_4	0.0113	0.0129	0.0132	0.0138	0.0141		0.0131		0.0116	0.0115	
Strontium thiosulfate	$\text{SrS}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$		2.5									
Strontium tungstate	SrWO_4			3.957×10^{-4}								
Sucrose	$\text{C}_{12}\text{H}_{22}\text{O}_{11}$	181.9	190.6	201.9	216.7	235.6	259.6	288.8	323.7	365.1	414.9	476.0
Sulfur dioxide	SO_2			9.4								

T

X

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Xenon</u>	<u>Xe</u>	24.1 ml		11.9 ml ²⁵			8.4 ml			7.12 ml		
<u>Xylose</u>	<u>C₅H₁₀O₅</u>			117								

Y

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Ytterbium(III) nitrate</u>	<u>Yb(NO₃)₃</u>			239								
<u>Ytterbium(III) sulfate</u>	<u>Yb₂(SO₄)₃</u>	44.2	37.5	38.4	22.2	17.2		10.4		6.4	5.8	4.7
<u>Yttrium(III) acetate</u>	<u>Y(C₂H₃O₂)₃·4H₂O</u>			9.03								
<u>Yttrium(III) bromate</u>	<u>Y(BrO₃)₃·9H₂O</u>			168								
<u>Yttrium(III) bromide</u>	<u>YBr₃</u>	63.9		75.1		87.3		101		116	123	
<u>Yttrium(III) chloride</u>	<u>YCl₃</u>	77.3	78.1	78.8	79.6	80.8						
<u>Yttrium(III) fluoride</u>	<u>YF₃</u>			0.005769								
<u>Yttrium(III) nitrate</u>	<u>Y(NO₃)₃</u>	93.1	106	123	143	163		200				
<u>Yttrium(III) sulfate</u>	<u>Y₂(SO₄)₃</u>	8.05	7.67	7.3	6.78	6.09		4.44		2.89	2.2	

Z

Substance	Formula	0 °C	10 °C	20 °C	30 °C	40 °C	50 °C	60 °C	70 °C	80 °C	90 °C	100 °C
<u>Zinc acetate</u>	<u>Zn(C₂H₃O₂)₂</u>			30								
<u>Zinc bromide</u>	<u>ZnBr₂</u>	389		446	528	591		618		645		672
<u>Zinc carbonate</u>	<u>ZnCO₃</u>			4.692×10 ⁻⁵								
<u>Zinc chlorate</u>	<u>Zn(ClO₃)₂</u>	145	152	209	223							
<u>Zinc chloride</u>	<u>ZnCl₂</u>	342	353	395	437	452		488		541		614
<u>Zinc cyanide</u>	<u>Zn(CN)₂</u>			0.058								
<u>Zinc fluoride</u>	<u>ZnF₂</u>			1.6								
<u>Zinc formate</u>	<u>Zn(HCO₂)₂</u>	3.7	4.3	6.1	7.4		11.8		21.2	28.8	38	
<u>Zinc iodate</u>	<u>Zn(IO₃)₂·2H₂O</u>			0.07749								
<u>Zinc iodide</u>	<u>ZnI₂</u>	430		432		445		467		490		510
<u>Zinc nitrate</u>	<u>Zn(NO₃)₂</u>	98			138	211						
<u>Zinc oxalate</u>	<u>ZnC₂O₄·2H₂O</u>			1.38×10 ⁻⁹								
<u>Zinc permanganate</u>	<u>Zn(MnO₄)₂</u>			33.3								
<u>Zinc sulfate</u>	<u>ZnSO₄</u>	41.6	47.2	53.8	61.3	70.5		75.4		71.1		60.5
<u>Zinc sulfite</u>	<u>ZnSO₃·2H₂O</u>			0.16								
<u>Zinc tartrate</u>	<u>ZnC₄H₄O₆</u>			0.022	0.041	0.06		0.104		0.59		
<u>Zirconium fluoride</u>	<u>ZrF₄</u>			1.32								
<u>Zirconium sulfate</u>	<u>Zr(SO₄)₂·4H₂O</u>			52.5								

External links

- Solubility Database (<http://srdata.nist.gov/solubility/index.aspx>) - International Union of Pure and Applied Chemistry / National Institute of Standards and Technology
- CRC Handbook of Chemistry and Physics (<http://www.hbcpnetbase.com>) - Online resource that includes solubility data

References

1. "Magnesium Sulfite" (https://srdata.nist.gov/solubility/IUPAC/SDS-26/SDS-26-pages_153.pdf) (PDF). *srdata.nist.gov*. Retrieved 2018-12-21.
2. Encyclopedia of the Alkaline Earth Compounds (Richard C. Ropp) page 86

- Chemcalc v4.0 - software that includes data on solubility
- Learning, Food resources (<https://web.archive.org/web/20061207132126/http://food.oregonstate.edu/learn/sugar.html>)
- Kaye and Laby Online (<http://www.kayelaby.npl.co.uk/>)
- ChemBioFinder.com (<https://web.archive.org/web/20141109053519/http://www.chemfinder.com/>) (registration required)

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