


☐

I'm not robot

  
reCAPTCHA

I am not robot!

First 50 elements and their symbols and atomic number

First 30 elements and their symbols and atomic number. What are the first 20 elements and their symbols and atomic number. What are the first 50 elements and their symbols.

List of the first 50 elements of the Mendeleiev table by atomic number, including the chemical symbol and the atomic weight. The list of positions can be printed by clicking on the Print button below. Mendeleev's painting for the group's home science chemistry is presented in the form of vertical columns, the number of which is 1 to 18. The elements of the group have very similar chemical properties, which are due to the number of valence electrons Present, that is to say the number of electrons in the external shell of the atom of an element in the table of Mendeleiev is given from the electronic configurations of the elements. According to the principle of Pauli, an orbital cannot fill more than two electrons. The first row of Mendeleev's painting includes only two elements: hydrogen and helium.

Table 4.1: Composition of Atoms of the First Eighteen Elements with Electron Distribution in Various Shells										
Name of Element	Symbol	Atomic Number	Number of	Number of	Number of	Distribution of Electrons				Valency
			Protons	Neutrons	Electrons	K	L	M	N	
Hydrogen	H	1	1	-	1	1	-	-	-	1
Helium	He	2	2	2	2	2	-	-	-	0
Lithium	Li	3	3	4	3	2	1	-	-	1
Beryllium	Be	4	4	5	4	2	2	-	-	2
Boron	B	5	5	6	5	2	3	-	-	3
Carbon	C	6	6	6	6	2	4	-	-	4
Nitrogen	N	7	7	7	7	2	5	-	-	3
Oxygen	O	8	8	8	8	2	6	-	-	2
Fluorine	F	9	9	10	9	2	7	-	-	1
Neon	Ne	10	10	10	10	2	8	-	-	0
Sodium	Na	11	11	12	11	2	8	1	-	1
Magnesium	Mg	12	12	12	12	2	8	2	-	2
Aluminium	Al	13	13	14	13	2	8	3	-	3
Silicon	Si	14	14	14	14	2	8	4	-	4
Phosphorus	P	15	15	16	15	2	8	5	-	3,5
Sulphur	S	16	16	16	16	2	8	6	-	2
Chlorine	Cl	17	17	18	17	2	8	7	-	1
Argon	Ar	18	18	22	18	2	8	8		0

The list of positions can be printed by clicking on the Print button below. Mendeleev's painting for the group's home science chemistry is presented in the form of vertical columns, the number of which is 1 to 18. The elements of the group have very similar chemical properties, which are due to the number of valence electrons Present, that is to say the number of electrons in the external shell of the atom of an element in the table of Mendeleiev is given from the electronic configurations of the elements. According to the principle of Pauli, an orbital cannot fill more than two electrons. The first row of Mendeleev's painting includes only two elements: hydrogen and helium. Because atoms have more electrons, there are more orbitals that can be filled, so more elements in the lines lower in the table. At the bottom of Mendeleiev's table are two lines generally separated from the main part of the table. These rows contain elements from a number of Lanthanides and Actinoids, generally 57 to 71 (Lanthane in Lutetia) and 89 to 103 (Actinium in Lawrence), respectively. There is no scientific basis for that. This is done only to make the table more compact. The periodic table, the full table of Mendeleiev, is in chemistry, a list of all the chemical elements in an increasing atomic number, that is to say Y., the total number of protons in the nucleus of an atom. In this arrangement of chemical elements, their properties follow a model known as "periodic law", in which the elements of the same column (group) have similar properties. The original discovery of Dmitri Ivanovitch Mendeleiev in the middle of the 19th century was invaluable for the development of chemistry. It was only in the second decade of the 20th century that the sequence of the elements of the periodic table was really recognized as their atomic numbering system.b' List of the first 50 elements of the periodic table by atomic number, including chemical symbol and atomic weight. You can print the list of items by clicking the "Print" button below. Home Science Chemistry Groups in the periodic table are displayed as vertical columns numbered 1 to 18. Elements in a group have very similar chemical properties, which are determined by the number of valence electrons present, which are the electrons in the outer shell of the atom. The arrangement of elements in the periodic table is determined by the electronic configuration of the elements. According to the Pauli principle, an orbital can be filled by no more than two electrons. The first row of the periodic table contains only two elements: hydrogen and helium.

Atomic number	Symbol	Electron configuration	Atomic number	Symbol	Electron configuration	Atomic number	Symbol	Electron configuration
1	H	1s¹	37	Rb	[Kr]5s¹	73	Ta	[Xe]4f¹⁴5d³6s²
2	He	1s²	38	Sr	[Kr]5s²	74	W	[Xe]4f¹⁴5d⁴6s²
3	Li	[He]2s¹	39	Y	[Kr]5s²4d¹	75	Re	[Xe]4f¹⁴5d⁵6s²
4	Be	[He]2s²	40	Zr	[Kr]5s²4d²	76	Os	[Xe]4f¹⁴5d⁶6s²
5	B	[He]2s²2p¹	41	Nb	[Kr]5s¹4d⁴	77	Ir	[Xe]4f¹⁴5d⁷6s²
6	C	[He]2s²2p²	42	Mo	[Kr]5s¹4d⁵	78	Pt	[Xe]4f¹⁴5d⁹6s¹
7	N	[He]2s²2p³	43	Tc	[Kr]5s³4d³	79	Au	[Xe]4f¹⁴5d¹⁰6s¹
8	O	[He]2s²2p⁴	44	Ru	[Kr]5s¹4d⁵	80	Hg	[Xe]4f¹⁴5d¹⁰6s²
9	F	[He]2s²2p⁵	45	Rh	[Kr]5s¹4d⁵	81	Tl	[Xe]4f¹⁴5d¹⁰6s²6p¹
10	Ne	[He]2s²2p⁶	46	Pd	[Kr]4d¹⁰	82	Pb	[Xe]4f¹⁴5d¹⁰6s²6p²
11	Na	[Ne]3s¹	47	Ag	[Kr]5s¹4d⁹	83	Bi	[Xe]4f¹⁴5d¹⁰6s²6p³
12	Mg	[Ne]3s²	48	Cd	[Kr]5s²4d¹⁰	84	Po	[Xe]4f¹⁴5d¹⁰6s²6p⁴
13	Al	[Ne]3s²3p¹	49	In	[Kr]5s²4d¹⁰5p²	85	At	[Xe]4f¹⁴5d¹⁰6s²6p⁵
14	Si	[Ne]3s²3p²	50	Sn	[Kr]5s²4d¹⁰5p²	86	Rn	[Xe]4f¹⁴5d¹⁰6s²6p⁶
15	P	[Ne]3s²3p³	51	Sb	[Kr]5s²4d¹⁰5p³	87	Fr	[Rn]7s¹
16	S	[Ne]3s²3p⁴	52	Te	[Kr]5s²4d¹⁰5p⁴	88	Ra	[Rn]2f¹
17	Cl	[Ne]3s²3p⁵	53	I	[Kr]5s²4d¹⁰5p⁵	89	Ac	[Rn]7s²6d¹
18	Ar	[Ne]3s²3p⁶	54	Xe	[Kr]5s²4d¹⁰5p⁶	90	Th	[Rn]7s²6d²
19	K	[Ar]4s¹	55	Cs	[Xe]6s¹	91	Pa	[Rn]7s²5f²6d¹
20	Ca	[Ar]4s²	56	Ba	[Xe]6s²	92	U	[Rn]7s²5f³6d¹
21	Sc	[Ar]4s²3d¹	57	La	[Xe]6s²5d¹	93	Np	[Rn]7s²5f⁴6d¹
22	Ti	[Ar]4s²3d²	58	Ce	[Xe]6s²4f¹5d¹	94	Pu	[Rn]7s²5f⁶
23	V	[Ar]4s²3d³	59	Pr	[Xe]6s²4f³	95	Am	[Rn]7s²5f⁷
24	Cr	[Ar]4s¹3d⁵	60	Nd	[Xe]6s²4f⁴	96	Cm	[Rn]7s²5f⁷6d¹
25	Mn	[Ar]4s²3d⁵	61	Pm	[Xe]6s²4f⁵	97	Bk	[Rn]7s²5f⁸
26	Fe	[Ar]4s²3d⁶	62	Sm	[Xe]6s²4f⁶	98	Cf	[Rn]7s²5f¹⁰
27	Co	[Ar]4s²3d⁷	63	Eu	[Xe]6s²4f⁷	99	Es	[Rn]7s²5f¹¹
28	Ni	[Ar]4s²3d⁸	64	Gd	[Xe]6s²4f⁷5d¹	100	Fm	[Rn]7s²5f¹²
29	Cu	[Ar]4s¹3d¹⁰	65	Tb	[Xe]6s²4f⁹	101	Md	[Rn]7s²5f¹³
30	Zn	[Ar]4s²3d¹⁰	66	Dy	[Xe]6s²4f¹⁰	102	No	[Rn]7s²5f¹⁴
31	Ga	[Ar]4s²3d¹⁰4p¹	67	Ho	[Xe]6s²4f¹¹	103	Lr	[Rn]7s²5f¹⁴6d¹
32	Ge	[Ar]4s²3d¹⁰4p²	68	Er	[Xe]6s²4f¹²	104	Rf	[Rn]7s²5f¹⁴6d²
33	As	[Ar]4s²3d¹⁰4p³	69	Tm	[Xe]6s²4f¹³	105	Db	[Rn]7s²5f¹⁴6d³
34	Se	[Ar]4s²3d¹⁰4p⁴	70	Yb	[Xe]6s²4f¹⁴	106	Sg	[Rn]7s²5f¹⁴6d⁴
35	Br	[Ar]4s²3d¹⁰4p⁵	71	Lu	[Xe]6s²4f¹⁴5d¹	107	Bh	[Rn]7s²5f¹⁴6d⁵
36	Kr	[Ar]4s²3d¹⁰4p⁶	72	Hf	[Xe]6s²4f¹⁴5d²	108	Hs	[Rn]7s²5f¹⁴6d⁶
						109	Mt	[Rn]7s²5f¹⁴6d⁷
						110	Dt	[Rn]7s²5f¹⁴6d⁸
						111	Rg	[Rn]7s²5f¹⁴6d⁹

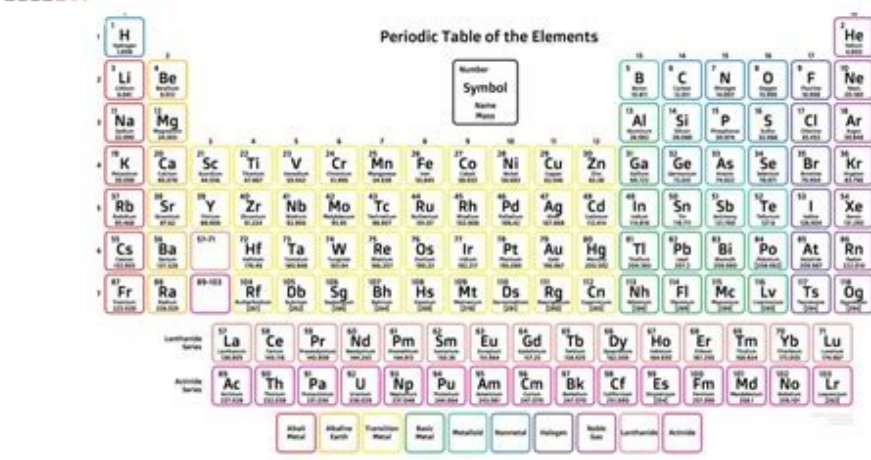
At the bottom of Mendeleiev's table are two lines generally separated from the main part of the table. These rows contain elements from a number of Lanthanides and Actinoids, generally 57 to 71 (Lanthane in Lutetia) and 89 to 103 (Actinium in Lawrence), respectively. There is no scientific basis for that. This is done only to make the table more compact. The periodic table, the full table of Mendeleiev, is in chemistry, a list of all the chemical elements in an increasing atomic number, that is to say Y., the total number of protons in the nucleus of an atom.

Composition of the atoms of the first eighteen elements :-										
Name of element	Symbol	Atomic Number	Number of Protons	Number of Neutrons	Number of Electrons	Distribution Of Electrons K L M N				Valency
Hydrogen	H	1	1	-	1	1	-	-	-	1
Helium	He	2	2	2	2	2	-	-	-	0
Lithium	Li	3	3	4	3	2	1	-	-	1
Beryllium	Be	4	4	5	4	2	2	-	-	2
Boron	B	5	5	6	5	2	3	-	-	3
Carbon	C	6	6	6	6	2	4	-	-	4
Nitrogen	N	7	7	7	7	2	5	-	-	3
Oxygen	O	8	8	8	8	2	6	-	-	2
Fluorine	F	9	9	10	9	2	7	-	-	1
Neon	Ne	10	10	10	10	2	8	-	-	0
Sodium	Na	11	11	12	11	2	8	1	-	1
Magnesium	Mg	12	12	12	12	2	8	2	-	2
Aluminium	Al	13	13	14	13	2	8	3	-	3
Silicon	Si	14	14	14	14	2	8	4	-	4
Phosphorus	P	15	15	16	15	2	8	5	-	3,5
Sulphur	S	16	16	16	16	2	8	6	-	2
Chlorine	Cl	17	17	18	17	2	8	7	-	1
Argon	Ar	18	18	22	18	2	8	8	-	0

The first row of Mendeleev's painting includes only two elements: hydrogen and helium. Because atoms have more electrons, there are more orbitals that can be filled, so more elements in the lines lower in the table. At the bottom of Mendeleiev's table are two lines generally separated from the main part of the table. These rows contain elements from a number of Lanthanides and Actinoids, generally 57 to 71 (Lanthane in Lutetia) and 89 to 103 (Actinium in Lawrence), respectively. There is no scientific basis for that. This is done only to make the table more compact. The periodic table, the full table of Mendeleiev, is in chemistry, a list of all the chemical elements in an increasing atomic number, that is to say Y., the total number of protons in the nucleus of an atom. In this arrangement of chemical elements, their properties follow a model known as "periodic law", in which the elements of the same column (group) have similar properties. The original discovery of Dmitri Ivanovitch Mendeleiev in the middle of the 19th century was invaluable for the development of chemistry. It was only in the second decade of the 20th century that the sequence of the elements of the periodic table was really recognized as their atomic numbering system.b' List of the first 50 elements of the periodic table by atomic number, including chemical symbol and atomic weight. You can print the list of items by clicking the "Print" button below. Home Science Chemistry Groups in the periodic table are displayed as vertical columns numbered 1 to 18. Elements in a group have very similar chemical properties, which are determined by the number of valence electrons present, which are the electrons in the outer shell of the atom. The arrangement of elements in the periodic table is determined by the electronic configuration of the elements. According to the Pauli principle, an orbital can be filled by no more than two electrons. The first row of the periodic table contains only two elements: hydrogen and helium. Because atoms have more electrons, they have more orbitals available to be filled, so rows further down the table contain more elements. At the end of the periodic table there are two rows that are usually separated from the main part of the periodic table. These series contain elements from the lanthanide and actinide series, typically 57 to 71 (lanthanum to lutetium) and 89 to 103 (actinium to lawrentium), respectively. There is no scientific basis for this. This is simply done to make the table more compact. Periodic table, complete periodic table, in chemistry, a table of all chemical elements arranged in increasing order of atomic numberxe2x80x94i.e. HOUR. The initial discovery of Dmitri Ivanovich Mendeleev in the mid-19th century was invaluable for the development of chemistry.



# Atomic Mass of Elements



The first row of Mendeleev's painting includes only two elements: hydrogen and helium. Because atoms have more electrons, there are more orbitals that can be filled, so more elements in the lines lower in the table. At the bottom of Mendeleiev's table are two lines generally separated from the main part of the table. These rows contain elements from a number of Lanthanides and Actinoids, generally 57 to 71 (Lanthane in Lutetia) and 89 to 103 (Actinium in Lawrence), respectively.

Element	Symbol	Z	Element	Symbol	Z
Hydrogen	H	1	Nickel	Ni	28
Helium	He	2	Copper	Cu	29
Carbon	C	6	Zinc	Zn	30
Nitrogen	N	7	Bromine	Br	35
Oxygen	O	8	Iodine	I	53
Fluorine	F	9	Barium	Ba	56
Sodium	Na	11	Tungsten	W	74
Magnesium	Mg	12	Platinum	Pt	78
Aluminium	Al	13	Gold	Au	79
Silicon	Si	14	Mercury	Hg	80
Phosphorus	P	15	Lead	Pb	82
Sulfur	S	16	Bismuth	Bi	83
Chlorine	Cl	17	Radium	Ra	88
Potassium	K	19	Thorium	Th	90
Calcium	Ca	20	Uranium	U	92
Scandium	Sc	21	Iron	Fe	26
Manganese	Mn	25	Cobalt	Co	27

At the bottom of Mendeleiev's table are two lines generally separated from the main part of the table. These rows contain elements from a number of Lanthanides and Actinoids, generally 57 to 71 (Lanthane in Lutetia) and 89 to 103 (Actinium in Lawrence), respectively. There is no scientific basis for that. This is done only to make the table more compact. The periodic table, the full table of Mendeleiev, is in chemistry, a list of all the chemical elements in an increasing atomic number, that is to say Y., the total number of protons in the nucleus of an atom.

In this arrangement of chemical elements, their properties follow a model known as "periodic law", in which the elements of the same column (group) have similar properties. The original discovery of Dmitri Ivanovitch Mendeleiev in the middle of the 19th century was invaluable for the development of chemistry. It was only in the second decade of the 20th century that the sequence of the elements of the periodic table was really recognized as their atomic numbering system.h' List of the first 50 elements of the periodic table by atomic number, including chemical symbol and atomic weight. You can print the list of items by clicking the "Print" button below. Home Science Chemistry Groups in the periodic table are displayed as vertical columns numbered 1 to 18. Elements in a group have very similar chemical properties, which are determined by the number of valence electrons present, which are the electrons in the outer shell of the atom. The arrangement of elements in the periodic table is determined by the electronic configuration of the elements. According to the Pauli principle, an orbital can be filled by no more than two electrons. The first row of the periodic table contains only two elements: hydrogen and helium. Because atoms have more electrons, they have more orbitals available to be filled, so rows further down the table contain more elements. At the end of the periodic table there are two rows that are usually separated from the main part of the periodic table. These series contain elements from the lanthanide and actinide series, typically 57 to 71 (lanthanum to lutetium) and 89 to 103 (actinium to lawrentium), respectively. There is no scientific basis for this. This is simply done to make the table more compact. Periodic table, complete periodic table, in chemistry, a table of all chemical elements arranged in increasing order of atomic number\xe2\x80\x94i.e. HOUR, the total number of protons in the nucleus of an atom. When chemical elements are arranged in this way, their properties exhibit a repeating pattern, called a periodic pattern, in which elements in the same column (group) have similar properties. The initial discovery of Dmitri Ivanovich Mendeleev in the mid-19th century was invaluable for the development of chemistry.

It was not until the second decade of the 20th century that the order of the elements in the periodic table was truly recognized.In this article, in the early 19th century, the work of rapid analytical chemistry of various chemicals-a breakdown of various chemicals and thus gathered much knowledge about the chemical and physical properties of elements and relationships. Due to the rapid development of chemistry, the development of chemistry required rapid classification, since the classification of chemistry is based not only on systematic chemical literature, but also on laboratory art, which transfers chemists from one generation of chemists to 'other. The connection was easier to see between relationships than elements; It just so happens that element classification has been behind relationship classification for many years. Indeed, chemists have not reached general agreement on the classification of elements nearly half a century after the widespread use of classification systems. Facts you should know: periodic kviz table J.W. 1817. Döbeiner showed that the page weight, i.e. the atomic mass, is half the calcium and Barrya burden, and a few years later showed that there was another triad (chlorine, bromine and iodine), halogens] and lithium, sodium and potassium (alkali metals)]. J.-B.-A. Dumas, L. Gmelin, E. Lenssen, Max von Pettenkofer, and J. P. Cooke expanded Dabereiner's suggestions in 1827-1858. He shows that a similar relationship lasts over three elements when fluorine was added to halogens and magnesium. The metallic alkali metal, while oxygen, sulfur, selenium and telide were attributed to one family and nitrogen, phosphorus, arsenic, antimony and bismuth were as another family of elements. Subsequently, attempts were made to demonstrate that the atomic mass of elements can be expressed in an arithmetic function. A.-E.-B. De Chancortois offered the classificationForecasts in the light of modern knowledge. Get a premium British subscription and access the exclusive content. Register for 1864, J.A.R. Newland proposed to classify the elements in the growing order of atomic weight, dividing the elements according to the number of unitary atoms and in seven groups made up of nitrogen and oxygen. These reports were called octaves by analogy to the seven intervals of the musical scale. In 1869, due to a complete correlation between the properties of the elements and the atomic weight, by paying particular attention to the value (that is to say the number of individual links of which an element can be established), Mendeleev proposed the periodic law, according to which the elements are arranged according to the size of the atomic weight, showing regular changes in the properties. Lothar Meier had come independently of a similar conclusion, which was published after the publication of Mendeleev's article. Allow JavaScript to use the Pubchem website. Website.