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## First 50 elements and their symbols and atomic number

First 30 elements and their symbols and atomic number. What are the first 20 elements and their symbols and atomic number. What are the first 50 elements and their symbols.

List of the first 50 elements of the Mendeleev table by atomic number, including the chemical symbol and the atomic weight. The list of positions can be printed by clicking on the Print button below. Mendeleev's painting for the group's home science chemistry is presented in the form of vertical columns, the number of which is 1 to 18. The elements of the group have very similar chemical properties, which are due to the number of valence electrons Present, that is to say the number of electrons in the external shell of the atom of an element in the table of Mendeleev is given from the electronic configurations of the elements. According to the principle of Pauli, an orbital cannot fill more than two electrons. The first row of Mendeleev's painting includes only two elements: hydrogen and helium.

Table 4.1: Composition of Atoms of the First Eighteen Elements with Electron Distribution in Various Shells											
Name of Element	Symbol	Atomic Number	Number of Protons	Number of Neutrons	Number of Electrons	Distribution of Electrons	K	L	M	N	Vale-ny
Hydrogen	H	1	1	-	1	1	-	-	-	-	1
Helium	He	2	2	2	2	2	2	-	-	-	0
Lithium	Li	3	3	4	3	2	1	-	-	-	1
Beryllium	Be	4	4	5	4	2	2	-	-	-	2
Boron	B	5	5	6	5	2	3	-	-	-	3
Carbon	C	6	6	6	6	2	4	-	-	-	4
Nitrogen	N	7	7	7	7	2	5	-	-	-	3
Oxygen	O	8	8	8	8	2	6	-	-	-	2
Fluorine	F	9	9	10	9	2	7	-	-	-	1
Neon	Ne	10	10	10	10	2	8	-	-	-	0
Sodium	Na	11	11	12	11	2	8	1	-	-	1
Magnesium	Mg	12	12	12	12	2	8	2	-	-	2
Aluminum	Al	13	13	14	13	2	8	3	-	-	3
Silicon	Si	14	14	14	14	2	8	4	-	-	4
Phosphorus	P	15	15	16	15	2	8	5	-	-	3.5
Sulphur	S	16	16	16	16	2	8	6	-	-	2
Chlorine	Cl	17	17	18	17	2	8	7	-	-	1
Argon	Ar	18	18	22	18	2	8	8	-	-	0

The list of positions can be printed by clicking on the Print button below. Mendeleev's painting for the group's home science chemistry is presented in the form of vertical columns, the number of which is 1 to 18. The elements of the group have very similar chemical properties, which are due to the number of valence electrons Present, that is to say the number of electrons in the external shell of the atom of an element in the table of Mendeleev is given from the electronic configurations of the elements. According to the principle of Pauli, an orbital cannot fill more than two electrons. The first row of Mendeleev's painting includes only two elements: hydrogen and helium. Because atoms have more electrons, there are more orbitals that can be filled, so more elements in the lines lower in the table. At the bottom of Mendeleev's table are two lines generally separated from the main part of the table. These rows contain elements from a number of Lanthanides and Actinoids, generally 57 to 71 (Lanthane in Lutetia) and 89 to 103 (Actinium in Lawrence), respectively. There is no scientific basis for that. This is done only to make the table more compact. The periodic table, the full table of Mendeleev, is in chemistry, a list of all the chemical elements in an increasing atomic number, that is to say Y., the total number of protons in the nucleus of an atom. In this arrangement of chemical elements, their properties follow a model known as "periodic law", in which the elements of the same column (group) have similar properties. The original discovery of Dmitri Ivanovich Mendeleev in the middle of the 19th century was invaluable for the development of chemistry. It was only in the second decade of the 20th century that the sequence of the elements of the periodic table was really recognized as their atomic numbering system.b List of the first 50 elements of the periodic table by atomic number, including chemical symbol and atomic weight. You can print the list of items by clicking the "Print" button below. Home Science Chemistry Groups in the periodic table are displayed as vertical columns numbered 1 to 18. Elements in a group have very similar chemical properties, which are determined by the number of valence electrons present, which are the electrons in the outer shell of the atom. The arrangement of elements in the periodic table is determined by the electronic configuration of the elements. According to the Pauli principle, an orbital can be filled by no more than two electrons. The first row of the periodic table contains only two elements: hydrogen and helium. The arrangement of elements in the periodic table is determined by the electronic configuration of the elements. According to the Pauli principle, an orbital can be filled by no more than two electrons. The first row of the periodic table contains only two elements: hydrogen and helium.

Atomic number	Symbol	Electron configuration	Atomic number	Symbol	Electron configuration	Atomic number	Symbol	Electron configuration
1	H	1s <sup>1</sup>	37	Rb	[Ne]3s <sup>1</sup>	73	Ta	[Ar]3d <sup>5</sup> 4s <sup>2</sup>
2	He	1s <sup>2</sup>	38	Sr	[Ar]3s <sup>2</sup>	74	W	[Ar]3d <sup>5</sup> 4s <sup>2</sup>
3	Li	1s <sup>2</sup> 2s <sup>1</sup>	39	Y	[Ar]3s <sup>2</sup> 4s <sup>1</sup>	75	Re	[Ar]3d <sup>5</sup> 4s <sup>2</sup>
4	Be	1s <sup>2</sup> 2s <sup>2</sup>	40	Zr	[Ar]3s <sup>2</sup> 4s <sup>2</sup>	76	Os	[Ar]3d <sup>5</sup> 4s <sup>2</sup>
5	B	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>1</sup>	41	Nb	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 4p <sup>1</sup>	77	Ir	[Ar]3d <sup>5</sup> 4s <sup>2</sup>
6	C	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>2</sup>	42	Mo	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 4p <sup>2</sup>	78	Pt	[Ar]3d <sup>5</sup> 4s <sup>2</sup>
7	N	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>3</sup>	43	Tc	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 4p <sup>3</sup>	79	Au	[Ar]3d <sup>5</sup> 4s <sup>2</sup>
8	O	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>4</sup>	44	Ru	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 4p <sup>4</sup>	80	Hg	[Ar]3d <sup>5</sup> 4s <sup>2</sup>
9	F	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>5</sup>	45	Rb	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 4p <sup>5</sup>	81	Tl	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>5</sup>
10	Ne	1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup>	46	Pb	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 4p <sup>6</sup>	82	Pb	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
11	Na	[Ne]3s <sup>1</sup>	47	Ag	[Ar]3s <sup>2</sup> 4s <sup>1</sup>	83	Bi	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
12	Mg	[Ne]3s <sup>2</sup>	48	Cd	[Ar]3s <sup>2</sup> 4s <sup>2</sup>	84	Po	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
13	Al	[Ne]3s <sup>2</sup> 3p <sup>1</sup>	49	In	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3p <sup>1</sup>	85	At	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
14	Si	[Ne]3s <sup>2</sup> 3p <sup>2</sup>	50	Sn	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3p <sup>2</sup>	86	Rh	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
15	P	[Ne]3s <sup>2</sup> 3p <sup>3</sup>	51	Te	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3p <sup>3</sup>	87	Ir	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
16	S	[Ne]3s <sup>2</sup> 3p <sup>4</sup>	52	Se	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3p <sup>4</sup>	88	Rs	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
17	Cl	[Ne]3s <sup>2</sup> 3p <sup>5</sup>	53	I	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3p <sup>5</sup>	89	Ac	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
18	Ar	[Ne]3s <sup>2</sup> 3p <sup>6</sup>	54	Xe	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3p <sup>6</sup>	90	Th	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
19	K	[Ar]3s <sup>1</sup>	55	Cs	[Ar]3s <sup>2</sup>	91	Pa	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
20	Ca	[Ar]3s <sup>2</sup>	56	Ba	[Ar]3s <sup>2</sup> 4s <sup>2</sup>	92	U	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
21	Sc	[Ar]3s <sup>2</sup> 3d <sup>1</sup>	57	La	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>1</sup>	93	Np	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
22	Ti	[Ar]3s <sup>2</sup> 3d <sup>2</sup>	58	Ce	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>2</sup>	94	Pu	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
23	V	[Ar]3s <sup>2</sup> 3d <sup>3</sup>	59	Pr	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>3</sup>	95	Am	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
24	Cr	[Ar]3s <sup>2</sup> 3d <sup>4</sup>	60	Nd	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>4</sup>	96	Cm	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
25	Mn	[Ar]3s <sup>2</sup> 3d <sup>5</sup>	61	Pm	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>5</sup>	97	Bk	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
26	Fe	[Ar]3s <sup>2</sup> 3d <sup>6</sup>	62	Sr	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>6</sup>	98	Cf	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
27	Co	[Ar]3s <sup>2</sup> 3d <sup>7</sup>	63	Eu	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>7</sup>	99	Es	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
28	Ni	[Ar]3s <sup>2</sup> 3d <sup>8</sup>	64	Gd	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>8</sup>	100	Fm	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
29	Cu	[Ar]3s <sup>2</sup> 3d <sup>10</sup>	65	Tb	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>10</sup>	101	Md	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
30	Zn	[Ar]3s <sup>2</sup> 3d <sup>10</sup> 4s <sup>1</sup>	66	Dy	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>10</sup> 4s <sup>1</sup>	102	No	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
31	Ga	[Ar]3s <sup>2</sup> 3d <sup>10</sup> 4s <sup>2</sup>	67	Ho	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>10</sup> 4s <sup>2</sup>	103	Lr	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
32	Ge	[Ar]3s <sup>2</sup> 3d <sup>10</sup> 4s <sup>2</sup> 4p <sup>1</sup>	68	Er	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>10</sup> 4s <sup>2</sup> 4p <sup>1</sup>	104	Rf	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
33	As	[Ar]3s <sup>2</sup> 3d <sup>10</sup> 4s <sup>2</sup> 4p <sup>2</sup>	69	Tm	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>10</sup> 4s <sup>2</sup> 4p <sup>2</sup>	105	Db	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
34	Se	[Ar]3s <sup>2</sup> 3d <sup>10</sup> 4s <sup>2</sup> 4p <sup>3</sup>	70	Vb	[Ar]3s <sup>2</sup> 4s <sup>2</sup> 3d <sup>10</sup> 4s <sup>2</sup> 4p <sup>3</sup>	106	Sg	[Ar]3d <sup>5</sup> 4s <sup>2</sup> 4p <sup>6</sup>
35	Br	[Ar]3s <sup>2</sup> 3d <sup>10</sup> 4s <sup>2</sup> 4p <sup>4</sup>	71	Lu	[Ar]3s <sup>2</sup>			

Periodic Table of the Elements																	
Element	Symbol	Z	Element	Symbol	Z	Element	Symbol	Z	Element	Symbol	Z	Element	Symbol	Z	Element	Symbol	Z
Hydrogen	H	1	Nickel	Ni	28												
Helium	He	2	Copper	Cu	29												
Carbon	C	6	Zinc	Zn	30												
Nitrogen	N	7	Boron	B	35												
Oxygen	O	8	Iodine	I	53												
Fluorine	F	9	Barium	Ba	56												
Sodium	Na	11	Tungsten	W	74												
Magnesium	Mg	12	Platinum	Pt	79												
Aluminum	Al	13	Gold	Au	79												
Silicon	Si	14	Mercury	Hg	80												
Phosphorus	P	15	Lead	Pb	82												
Sulfur	S	16	Bismuth	Bi	83												
Chlorine	Cl	17	Radium	Ra	88												
Potassium	K	19	Thorium	Th	90												
Calcium	Ca	20	Uranium	U	92												
Sodium	Sc	21	Iron	Fe	26												
Manganese	Mn	25	Cobalt	Co	27												

The first row of Mendeleev's painting includes only two elements: hydrogen and helium. Because atoms have more electrons, there are more orbitals that can be filled, so more elements in the lines lower in the table. At the bottom of Mendeleev's table are two lines generally separated from the main part of the table. These rows contain elements from a number of Lanthanides and Actinoids, generally 57 to 71 (Lanthane in Lutetia) and 89 to 103 (Actinium in Lawrence), respectively.

Element	Symbol	Z	Element	Symbol	Z
Hydrogen	H	1	Nickel	Ni	28
Helium	He	2	Copper	Cu	29
Carbon	C	6	Zinc	Zn	30
Nitrogen	N	7	Boron	B	35
Oxygen	O	8	Iodine	I	53
Fluorine	F	9	Barium	Ba	56
Sodium	Na	11	Tungsten	W	74
Magnesium	Mg	12	Platinum	Pt	79
Aluminum	Al	13	Gold	Au	79
Silicon	Si	14	Mercury	Hg	80
Phosphorus	P	15	Lead	Pb	82
Sulfur	S	16	Bismuth	Bi	83
Chlorine	Cl	17	Radium	Ra	88
Potassium	K	19	Thorium	Th	90
Calcium	Ca	20	Uranium	U	92
Sodium	Sc	21	Iron	Fe	26
Manganese	Mn	25	Cobalt	Co	27

At the bottom of Mendeleev's table are two lines generally separated from the main part of the table. These rows contain elements from a number of Lanthanides and Actinoids, generally 57 to 71 (Lanthane in Lutetia) and 89 to 103 (Actinium in Lawrence), respectively. There is no scientific basis for that. This is done only to make the table more compact. The periodic table, the full table of Mendeleev, is in chemistry, a list of all the chemical elements in an increasing atomic number, that is to say Y, the total number of protons in the nucleus of an atom.

The original discovery of Dmitri Ivanovich Mendeleev in the middle of the 19th century was invaluable for the development of chemistry. It was only in the second decade of the 20th century that the sequence of the elements of the periodic table was really recognized as their atomic numbering system.b' List of the first 50 elements of the periodic table by atomic number, including chemical symbol and atomic weight. You can print the list of items by clicking the "Print" button below. Home Science Chemistry Groups in the periodic table are displayed as vertical columns numbered 1 to 18. Elements in a group have very similar chemical properties, which are determined by the number of valence electrons present, which are the electrons in the outer shell of the atom. The arrangement of elements in the periodic table is determined by the electronic configuration of the elements. According to the Pauli principle, an orbital can be filled by no more than two electrons. The first row of the periodic table contains only two elements: hydrogen and helium. Because atoms have more electrons, they have more orbitals available to be filled, so rows further down the table contain more elements. At the end of the periodic table there are two rows that are usually separated from the main part of the periodic table. These series contain elements from the lanthanide and actinide series, typically 57 to 71 (lanthanum to lutetium) and 89 to 103 (actinium to lawrentium), respectively. There is no scientific basis for this. This is simply done to make the table more compact. Periodic table, complete periodic table, in chemistry, a table of all chemical elements arranged in increasing order of atomic numberxe2x80x94, i.e. HOUR, the total number of protons in the nucleus of an atom. When chemical elements are arranged in this way, their properties exhibit a repeating pattern, called a periodic pattern, in which elements in the same column (group) have similar properties. The initial discovery of Dmitri Ivanovich Mendeleev in the mid-19th century was invaluable for the development of chemistry.

It was not until the second decade of the 20th century that the order of the elements in the periodic table was truly recognized. In this article, in the early 19th century, the work of rapid analytical chemistry of various chemicals-a breakdown of various chemicals and thus gathered much knowledge about the chemical and physical properties of elements and relationships. Due to the rapid development of chemistry, the development of chemistry required rapid classification, since the classification of chemistry is based not only on systematic chemical literature, but also on laboratory art, which transfers chemists from one generation of chemists to 'other. The connection was easier to see between relationships than elements; It just so happens that element classification has been behind relationship classification for many years. Indeed, chemists have not reached general agreement on the classification of elements nearly half a century after the widespread use of classification systems. Facts you should know: periodic kviz table J.W. Döbeleiner showed that the page weight, i.e. the atomic mass, is half the calcium and Barry burden, and a few years later showed that there was another triad (chlorine, bromine and iodine), halogens and lithium, sodium and potassium [alkali metals]. J.-B.-A. Dumas, L. Gmelin, Max von Pettenkofer, and J. P. Cooke expanded Dabereiner's suggestions in 1827-1858. He shows that a similar relationship lasts over three elements when chlorine was added to halogens and magnesium. The metallic alkali metal, while oxygen, sulfur, selenium and telluride were attributed to one family and nitrogen, phosphorus, arsenic, antimony and bismuth were as another family of elements. Subsequently, attempts were made to demonstrate that the atomic mass of elements can be expressed in an arithmetic function. A.-E.-B. De Chancortier offered the classificationForecasts in the light of modern knowledge. Get a premium British subscription and access the exclusive content. Register for 1864, J.A.R. Newland proposed to classify the elements in the growing order of atomic weight, dividing the elements according to the number of unitary atoms and in seven groups made up of nitrogen and oxygen. These reports were called octaves by analogy to the seven intervals of the musical scale. In 1869, due to a complete correlation between the properties of the elements and the atomic weight, by paying particular attention to the value (that is to say the number of individual links of which an element can be established), Mendeleev proposed the periodic law, according to which the elements are arranged according to the size of the atomic weight, showing regular changes in the properties. Lothar Meier had come independently of a similar conclusion, which was published after the publication of Mendeleev's article. Allow JavaScript to use the Pubchem website. Website.