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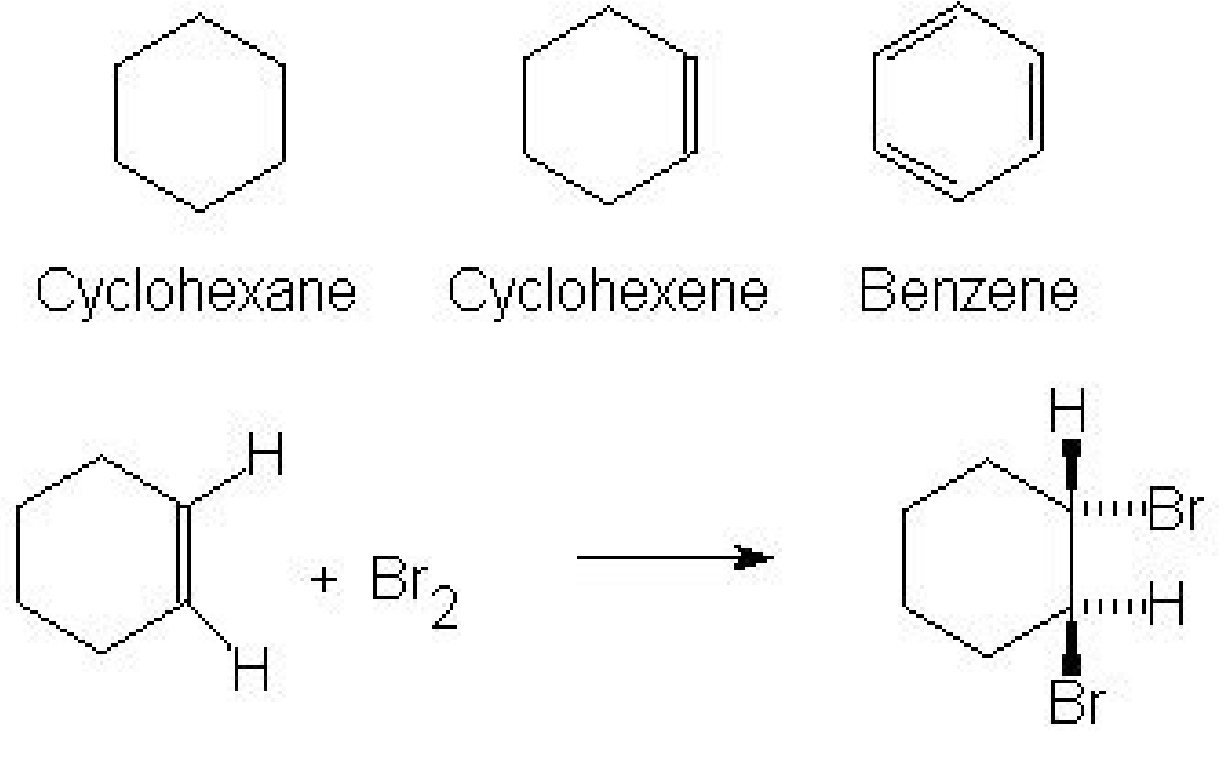

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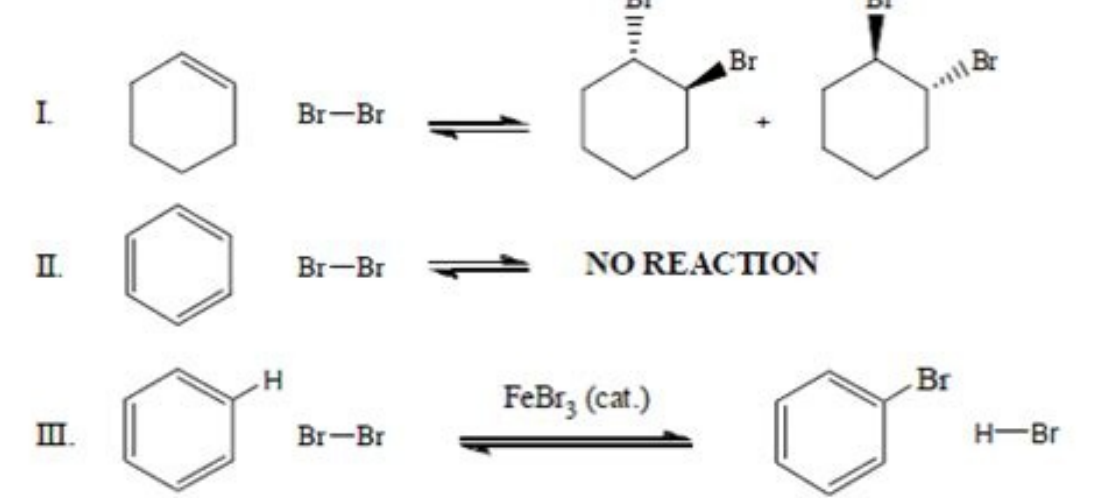
Cyclohexene and bromine chemical equation

Cyclohexene and bromine balanced equation. Does cyclohexene react with bromine. What happens when cyclohexene reacts with bromine.

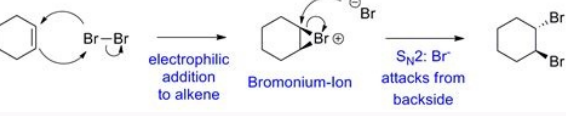
Home Subjects Math Science History Arts & Humanities Social Studies Engineering & Technology Business Other Resources Study Guides Leaderboard All Tags 7 Unanswered Random Tags Elements and Compounds Cyclohexene reacts with bromine in the same way and under the same conditions as any other alkene. 1,2-dibromocyclohexane is formed. The reaction is an example of electrophilic addition. How does cyclohexene react with bromine? Cyclohexene has no pi-unsaturation and is therefore not nucleophilic. It does not react with bromine unless energy in the form of light or heat is applied. In such a case a free-radical substitution reaction occurs. Cyclohexene is a typical alkene, and benzene and anisole are aromatic compounds. What is the product of the reaction between cyclohexane and one mole of bromine water br2 in the presence of UV light? The reaction between hexene and bromine in presence of light gives 3-bromocyclohexene. What happens when bromine solution is first added to cyclohexane? Bromine adds across the double bond of cyclohexene forming a clear solution of trans-1,2-Dibromocyclohexane. The cylinder containing cyclohexene remains colored.



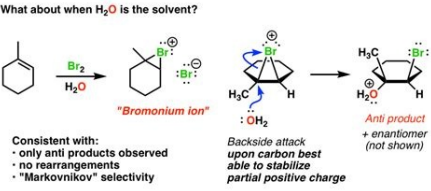
What is the chemical formula of cyclohexane? C6H12 Cyclohexane/Formula Does cyclohexene react with water?



Water, being one of the products, is insoluble in cyclohexene so forms a second layer. How do alkenes react with bromine water? Bromine water is an orange solution of bromine.



It becomes colourless when it is shaken with an alkene. This has the effect of ‘saturating’ the molecule, and will turn an alkene into an alkane. For example: C 2H 4 + H 2 → C 2H. What is the molecular formula of cyclohexene? C6H10 Cyclohexene/Formula CHEBI:36404 - cyclohexene Cyclohexene is a hydrocarbon with the formula C6H10. This cycloalkene is a colorless liquid with a sharp smell.



It is an intermediate in various industrial processes. Cyclohexene is not very stable upon long term storage with exposure to light and air because it forms peroxides. C6H10 + Br2 → C6H10Br2 The double bond within cyclohexene is opened and one bromine atom is added to each of the two carbon atoms that were originally double bonded. Both products are formed although 3-bromocyclohexene is the major product. Formation of major product: 3-bromocyclohexene Under UV light, %ce{Br2}\$ undergoes homolytic splitting to generate %ce{Br*}\$ radicals: %ce{Br2 ->[hv] 2Br*}\$ The formation of 3-bromocyclohexene is an example of substitution of alkanes, which require the free-radical mechanism: In the first step of the upper mechanism, which is also the rate-determining step, a stable allyl radical is generated, which is stabilized by resonance: As a result, the activation energy of the first step is significantly lowered. Formation of minor product: 1,2-dibromocyclohexane Individual bromine radicals are not electrophilic enough to attack the double bond in the cyclohexene, so the formation of 1,2-dibromocyclohexane requires the ions mechanism, typical for addition reactions (the lower mechanism in the following diagram). The first step in this mechanism is the rate-determining step. In this step, bromine is ionized, which requires a moderate amount of activation energy, albeit still much higher than the rate-determining step of the upper mechanism. Conclusion Therefore, the upper mechanism occurs at a much faster rate than the lower mechanism, which makes the major product 3-bromocyclohexene and the minor product 1,2-dibromocyclohexane. PS: Many people think that addition reaction is very fast.



It is only true in water, where the bromine ion is stabilized by solvation in water. Disclaimer: The mechanism I used for the addition reaction probably contradicts with your book. However, it does not matter. The point is that an ion is formed which makes the activation energy high. Bromine is highly toxic and reactive. Use care. Both cyclohexane and cyclohexene are very flammable. Keep away from flames. Cyclohexene and cyclohexane have an irritating odor. Dichloromethane is a cancer suspect agent. Bromine solution (5% in dichloromethane) Cyclohexane Cyclohexene 2 hydrometer cylinders Procedure Fill 1 cylinder 1/2 full with cyclohexane and fill the other 1/2 full with cyclohexene. Add bromine solution to both cylinders. Bromine adds across the double bond of cyclohexene forming a clear solution of trans-1,2-Dibromocyclohexane. The cylinder containing cyclohexane remains colored. Share Back to lecture demo index To schedule a demonstration, please login to the online lecture demonstration scheduler. Login with your netid in the form of "netid" Example: netid@umich.edu THE REACTION BETWEEN SYMMETRICAL ALKENES AND BROMINE This page gives you the facts and a simple, uncluttered mechanism for the electrophilic addition reactions between bromine (and the other halogens) and alkenes like ethene and cyclohexene. If you want the mechanisms explained to you in detail, there is a link at the bottom of the page. The facts Alkenes react in the cold with pure liquid bromine, or with a solution of bromine in an organic solvent like tetrachloromethane. The double bond breaks, and a bromine atom becomes attached to each carbon. The bromine loses its original red-brown colour to give a colourless liquid. In the case of the reaction with ethene, 1,2-dibromoethane is formed. This decolourisation of bromine is often used as a test for a carbon-carbon double bond. If an aqueous solution of bromine is used ("bromine water"), you get a mixture of products. The presence of the water complicates the mechanism beyond what is required by current UK A level (or equivalent) syllabuses. The other halogens, apart from fluorine, behave similarly. (Fluorine reacts explosively with all hydrocarbons - including alkenes - to give carbon and hydrogen fluoride.) If you are interested in the reaction with, say, chlorine, all you have to do is to replace Br by Cl in all the equations on this page. The mechanism for the reaction between ethene and bromine The reaction is an example of electrophilic addition. Cyclohexene Names Preferred IUPAC name Cyclohexene Other names Tetrahydrobenzene, 1,2,3,4-Tetrahydrobenzene, Benzenetetrahydride, Cyclohex-1-ene, Hexanaphthylene, UN 2256 Identifiers CAS Number 110-83-8 Y 3D model (JSmol) Interactive image Beilstein Reference 906737 ChEBI CHEBI:36404 Y ChEMBL ChEMBL16396 Y ChemSpider 7788 Y ECHA InfoCard 100.003.462 EC Number 203-807-8 Gmelin Reference 1659 PubChem CID 8079 RTECS number GW2500000 UNII 12L0P8F7GN Y CompTox Dashboard (EPA) DTXSID9038717 InChI InChI=1S/C6H10/c1-2-4-6-5-3-1/h1-2H,3-6H2 YKey: HGCIXCUEYOPUTN-UHFFFAOYSA-N YInChI=1/C6H10/c1-2-4-6-5-3-1/h1-2H,3-6H2Key: HGCIXCUEYOPUTN-UHFFFAOYSA-N SMILES C1CCC=CC1 Properties Chemical formula C6H10 Molar mass 82.143 g/mol Appearance colorless liquid Odor sweet Density 0.8110 g/cm3 Melting point −103.5 °C (−154.3 °F; 169.7 K) Boiling point 82.98 °C (181.36 °F; 356.13 K) Solubility in water slightly soluble in water Solubility miscible with organic solvents Vapor pressure 8.93 kPa (20 °C) 11.9 kPa (25 °C) Henry's lawconstant (kH) 0.022 mol·kg−1·bar−1 Magnetic susceptibility (χ) −57.5·10−6 cm3/mol Refractive index (nD) 1.4465 Hazards GHS labelling: Pictograms Signal word Danger Hazard statements H225, H302, H305, H311, H411 Precautionary statements P210, P233, P240, P241, P242, P243, P264, P270, P273, P280, P301+P310, P301+P312, P302+P352, P303+P361+P353, P312, P322, P330, P331, P361, P363, P370+P378, P391, P403+P235, P405, P501 NFPA 704 (fire diamond) 1 3 0 Flash point −12 °C (10 °F; 261 K) Autoignitiontemperature 244 °C (471 °F; 517 K) Explosive limits 0.8-5 % Lethal dose or concentration (LD, LC): LD50 (median dose) 1407 mg/kg (oral, rat) LCLo (lowest published) 13,196 ppm (mouse, 2 hr)[2] NIOSH (US health exposure limits): PEL (Permissible) TWA 300 ppm (1015 mg/m3)[1] REL (Recommended) TWA 300 ppm (1015 mg/m3)[1] IDLH (Immediate danger) 2000 ppm[1] Safety data sheet (SDS) External MSDS Except where otherwise noted, data are given for materials in their standard state (at 25 °C [77 °F], 100 kPa). Y verify (what is YN ?) Infobox references Chemical compound Cyclohexene is a hydrocarbon with the formula C6H10. This cycloalkene is a colorless liquid with a sharp smell. It is an intermediate in various industrial processes. Cyclohexene is not very stable upon long term storage with exposure to light and air because it forms peroxides. Production and uses Cyclohexene is produced by the partial hydrogenation of benzene, a process developed by the Asahi Chemical company.[3] In the laboratory, it can be prepared by dehydration of cyclohexanol.[4] Reactions and uses Benzene is converted to cyclohexylbenzene by acid-catalyzed alkylation with cyclohexene.[5] Cyclohexylbenzene is a precursor to both phenol and cyclohexanone.[6] Hydration of cyclohexene gives cyclohexanol, which can be dehydrogenated to give cyclohexanone, a precursor to caprolactam.[7] The oxidative cleavage of cyclohexene gives adipic acid. Hydrogen peroxide is used as the oxidant in the presence of a tungsten catalyst.[8] Bromination gives 1,2-dibromocyclohexane.[9] Structure Cyclohexene is most stable in a half-chair conformation,[10] unlike the preference for a chair form of cyclohexane. One basis for the cyclohexane conformational preference for a chair is that it allows each bond of the ring to adopt a staggered conformation. For cyclohexene, however, the alkene is planar, equivalent to an eclipsed conformation at that bond. See also Diels-Alder reaction Cyclohexa-1,3-diene Cyclohexa-1,4-diene References ^ a b c NIOSH Pocket Guide to Chemical Hazards. "#0167". National Institute for Occupational Safety and Health (NIOSH). ^ "Cyclohexene". Immediately Dangerous to Life or Health Concentrations (IDLH). National Institute for Occupational Safety and Health (NIOSH). ^ US 9771313, Narisawa, Naoki & Tanaka, Katsutoshi, "Cyclohexanol, method for producing cyclohexanol, and method for producing adipic acid", published 26 Sep 2017 ^ G. H. Coleman, H. F. Johnstone (1925). "Cyclohexene". Organic Syntheses. 5: 33. doi:10.15227/orgsyn.005.0033. ^ B. B. Corson, V. N. Ipatieff (1939). "Cyclohexylbenzene". Organic Syntheses. 19: 36. doi:10.15227/orgsyn.019.0036. ^ Plotkin, Jeffrey S. (2016-03-21). "What's New in Phenol Production?". American Chemical Society. Archived from the original on 2019-10-27. Retrieved 2018-01-02. ^ Musser, Michael T. (2005). "Cyclohexanol and Cyclohexanone". Ullmann's Encyclopedia of Industrial Chemistry. Weinheim: Wiley-VCH. doi:10.1002/14356007.a08.217. ^ Reed, Scott M.; Hutchison, James E. (2000). "Green Chemistry in the Organic Teaching Laboratory: An Environmentally Benign Synthesis of Adipic Acid". J. Chem. Educ. 77 (12): 1627–1629. Bibcode:2000JChEd..77.1627R. doi:10.1021/ed077p1627. ^ H. R. Snyder, L. A. Brooks (1932). "1,2-Dibromocyclohexane". Organic Syntheses. 12: 26. doi:10.15227/orgsyn.012.0026. ^ Jensen, Frederick R.; Bushweller, C. Hackett (1969). "Conformational preferences and interconversion barriers in cyclohexene and derivatives". J. Am. Chem. Soc. 91 (21): 5774–5782. doi:10.1021/ja01049a013.

External links from International Chemical Safety Card 1054 NIOSH Pocket Guide to Chemical Hazards. "#0167". National Institute for Occupational Safety and Health (NIOSH). Material Safety Data Sheet for Cyclohexene Safety MSDS data Reaction of Cyclohexene with Bromine and Potassium Permanganate Cyclohexene synthesis Data sheet at inchem.org Retrieved from " 2Chemical compound Cyclohexa-1,3-diene Names Preferred IUPAC name Cyclohexa-1,3-diene Other names 1,3-Cyclohexadiene, 1,2-Dihydrobenzene, 1,3-CHD, Benzene Identifiers CAS Number 592-57-4 Y 3D model (JSmol) Interactive image Beilstein Reference 506024 ChEBI CHEBI:37610 Y ChemSpider 11117 Y ECHA InfoCard 100.008.878 EC Number 209-764-1 Gmelin Reference 1657 PubChem CID 11605 RTECS number GU4702350 UNII JV5W0EG5BP Y UN number 1993 CompTox Dashboard (EPA) DTXSID30862259 InChI InChI=1S/C6H8/c1-2-4-6-5-3-1/h1-4H,5-6H2 YKey: MGNZXYYWBUKAII-UHFFFAOYSA-N YInChI=1/C6H8/c1-2-4-6-5-3-1/h1-4H,5-6H2Key: MGNZXYYWBUKAII-UHFFFAOYAH SMILES C1=C/C=C/CC/1 Properties Chemical formula C6H8 Molar mass 80.13 g/mol Appearance Colorless liquid Density 0.841 g/cm3 Melting point −98 °C (−144 °F; 175 K) Boiling point 80 °C (176 °F; 353 K) Magnetic susceptibility (χ) −48.6·10−6 cm3/mol Hazards GHS labelling: Pictograms Signal word Danger Hazard statements H225, H335 Precautionary statements P210, P233, P240, P241, P242, P243, P261, P271, P280, P303+P361+P353, P304+P340, P312, P370+P378, P403+P233, P403+P235, P405, P501 Flash point 26 °C (79 °F; 299 K) c.c. Except where otherwise noted, data are given for materials in their standard state (at 25 °C [77 °F], 100 kPa).

Y verify (what is YN ?) Infobox references Chemical compound Cyclohexa-1,3-diene (also known as Benzene) is an organic compound with the formula (C2H4)(CH)4. It is a colorless, flammable liquid. Its refractive index is 1.475 (20 °C, D). A naturally occurring derivative of cyclohexa-1,3-diene is terpinene, a component of pine oil. Synthesis Cyclohexadiene is prepared by the dehydrobromination of 1,2-dibromocyclohexane:[1] (CH2)4(CHBr)2 + 2 NaH → (CH2)2(CH)4 + 2 NaBr + 2 H2 Reactions Useful reactions of this diene are cycloadditions, such as the Diels-Alder reaction.[2] Conversion of cyclohexa-1,3-diene to benzene + hydrogen is exothermic by about 25 kJ/mol in the gas phase. [3][4] cyclohexane → cyclohexa-1,3-diene + 2 H2 (ΔH = +231.5 kJ/mol; endothermic) cyclohexane → benzene + 3 H2 (ΔH = +205 kJ/mol; endothermic) cyclohexa-1,3-diene → benzene + H2 (ΔH = −26.5 kJ/mol; exothermic) Compared with its isomer cyclohexa-1,4-diene, cyclohexa-1,3-diene is about 1.6 kJ/mol more stable.[5] Cyclohexadiene and its derivatives form metal-alkene complexes. Illustrative is [(C6H8)Fe(CO)3], an orange liquid. This complex reacts with hydride-abstracting reagents to give the cyclohexadienyl derivative [(C6H7)Fe(CO)3]+.[6] Cyclohexadienes react with ruthenium trichloride to give (Benzene)ruthenium dichloride dimer.[7] See also 1,4-Cyclohexadiene Cyclohexene References ^ Schaefer, John P.; Endres, Leland (1967). "1,3-Cyclohexadiene". *Organic Syntheses*. 47: 31. doi:10.15227/orgsyn.047.0031. ^ Sanjeeva Rao Guppi, George A. O'Doherty, "1,3-Cyclohexadiene" *Encyclopedia of Reagents for Organic Synthesis*, 2008 John Wiley & Sons. doi:10.1002/047084289X.rm00921 ^ US National Institute of Standards and Technology, NIST Chemistry WebBook 1,3-Cyclohexadiene Benzene ^ J. Sherman The heats of hydrogenation of unsaturated hydrocarbons Archived 2011-07-14 at the Wayback Machine *Journal of the American Oil Chemists' Society*; Volume 16, Number 2 / February, 1939 ^ NIST Chemistry WebBook 1,4-Cyclohexadiene ^ Pearson, Anthony J.; Sun, Huikai (2008). "Cyclohexadieneiron Tricarbonyl". *E-EROS Encyclopedia of Reagents for Organic Synthesis*. doi:10.1002/047084289X.rm00791. ISBN 978-0471936237. ^ Bennett, M. A.; Huang, T. N.; Matheson, T. W.; Smith, A. K. (1982). "16. (η6-Hexamethylbenzene)Ruthenium Complexes". (η6-Hexamethylbenzene)ruthenium Complexes. *Inorganic Syntheses*. Vol. 21. pp. 74–78. doi:10.1002/9780470132524.ch16. ISBN 9780470132524. Retrieved from " 3 Cyclohexa-1,4-diene Names Preferred IUPAC name Cyclohexa-1,4-diene[1] Other names 1,4-Cyclohexadiene[citation needed] 1,4-Dihydrobenzene[citation needed] Identifiers CAS Number 628-41-1 Y 3D model (JSmol) Interactive image Abbreviations 1,4-CHDN Beilstein Reference 1900733 ChEBI CHEBI:37611 Y ChemSpider 11838 Y ECHA InfoCard 100.010.040 EC Number 211-043-1 Gmelin Reference 1656 MeSH 1,4-cyclohexadiene PubChem CID 12343 UNII 0F8Z5909QZ Y UN number 3295 CompTox Dashboard (EPA) DTXSID0060854 InChI InChI=1S/C6H8/c1-2-4-6-5-3-1/h1-2,5-6H,3-4H2 YKey: UVJHQYIOXKWHFD-UHFFFAOYSA-N Y SMILES C1C=CCC=C1 Properties Chemical formula C6H8 Molar mass 80.130 g·mol−1 Appearance Colorless liquid Density 0.847 g cm−3 Melting point −50 °C; −58 °F; 223 K Boiling point 82 °C; 179 °F; 355 K Magnetic susceptibility (χ) −48.7·10−6 cm3/mol Refractive index (nD) 1.472 Thermochemistry Heat capacity (C) 142.2 J K−1 mol−1 Std molarentropy (S298) 189.37 J K−1 mol−1 Std enthalpy offormation (ΔfH298) 63.0-69.2 kJ mol−1 Std enthalpy ofcombustion (ΔcH298) -3573.5−3567.5 kJ mol−1 Hazards GHS labelling: Pictograms Signal word Danger Hazard statements H225, H340, H350, H373 Precautionary statements P201, P210, P308+P313 NFPA 704 (fire diamond) 2 3 0 Flash point −7 °C (19 °F; 266 K) Except where otherwise noted, data are given for materials in their standard state (at 25 °C [77 °F], 100 kPa).

Y verify (what is YN ?) Infobox references Chemical compound 1,4-Cyclohexadiene is an organic compound with the formula C6H8. It is a colourless, flammable liquid that is of academic interest as a prototype of a large class of related compounds called terpenoids, an example being γ-terpinene. An isomer of this compound is 1,3-cyclohexadiene. Synthesis and reactions In the laboratory, substituted 1,4-cyclohexadienes are synthesized by Birch reduction of related aromatic compounds using an alkali metal dissolved in liquid ammonia and a proton donor such as an alcohol. In this way, over reduction to the fully saturated ring is avoided, 1,4-Cyclohexadiene and its derivatives are easily aromatized, the driving force being the formation of an aromatic ring. The conversion to an aromatic system may be used to trigger other reactions, such as the Bergman cyclization.[2] References ^ "1,4-cyclohexadiene - Compound Summary". PubChem Compound. USA: National Center for Biotechnology Information. 27 March 2005. Identification and Related Records. Retrieved 12 October 2011. ^ John C. Walton, Fernando Portela-Cubillo "1,4-Cyclohexadiene" *Encyclopedia of Reagents for Organic Synthesis* 2007 John Wiley & Sons. doi:10.1002/047084289X.rm00806 External links The photochemistry of 1,4-cyclohexadiene in solution and in the gas phase NIST Chemistry WebBook Reaction thermochemistry data Retrieved from " 4This article needs additional citations for verification. Please help improve this article by adding citations to reliable sources. Unsourced material may be challenged and removed.Find sources: "Cyclohepta-1,4-diene" - news - newspapers - books - scholar · JSTOR (April 2022) (Learn how and when to remove this template message) Cyclohepta-1,4-diene Names Preferred IUPAC name Cyclohepta-1,4-diene Other names 1,4-Cycloheptadiene Identifiers CAS Number 7161-35-5 Y 3D model (JSmol) Interactive image Abbreviations 1,4-CHDN ChemSpider 122548 Y MeSH 1,4-cycloheptadiene PubChem CID 138950 UNII EFF73R8YGV Y UN number 3295 CompTox Dashboard (EPA) DTXSID00221859 InChI Key: HQGYGGZHWXFSI-UHFFFAOYSA-NInChI=1S/C7H10/c1-2-4-6-7-5-3-1/h1-2,5,7H,3-4,6H2 SMILES C1CC=CCC=C1 Properties Chemical formula C7H10 Molar mass 94.157 g·mol−1 Appearance Colorless liquid Density 0.84 g cm−3 Boiling point −150.445 °C; −238.801 °F; 122.705 K Refractive index (nD) 1.48 Hazards GHS labelling: Pictograms Signal word Danger Hazard statements H225, H340, H350, H373 Precautionary statements P201, P210, P308+P313 NFPA 704 (fire diamond) 2 3 0 Flash point 6.293 °C (43.327 °F; 279.443 K) Except where otherwise noted, data are given for materials in their standard state (at 25 °C [77 °F], 100 kPa). Infobox references Chemical compound 1,4-Cycloheptadiene is a highly flammable cycloalkene that occurs as a colorless clear liquid. It can form a yellow complex with palladium.[1] References ^ Rettig, Michael F.; Wing, Richard M. (June 1981). "X-ray crystallographic, chemical, and spectroscopic studies of the palladium dichloride complexes of cyclonona-1,5-diene, cycloocta-1,5-diene, cycloocta-1,4-diene, and cyclohepta-1,4-diene". *Journal of the American Chemical Society*. 103 (11): 2980–2986. doi:10.1021/ja00401a012. This article about a hydrocarbon is a stub. You can help Wikipedia by expanding it.vte Retrieved from " 5 1,3-Cycloheptadiene Names Preferred IUPAC name Cyclohepta-1,3-diene Other names 1,3-Cycloheptadiene Identifiers CAS Number 4054-38-0 Y 3D model (JSmol) Interactive image Abbreviations 1,3-CHDN Beilstein Reference 1900733 ChEBI CHEBI:37611 Y ChemSpider 1847604 Y ECHA InfoCard 100.021.603 EC Number 223-762-8 Gmelin Reference 1656 MeSH 1,3-cycloheptadiene PubChem CID 19969 UN number 3295 CompTox Dashboard (EPA) DTXSID40193566 SMILES C1CC=CC=CC1 Properties Chemical formula C7H10 Molar mass 94.157 g·mol−1 Appearance Colorless liquid Density 0.868 g cm−3 Melting point −110.40 °C; −166.72 °F; 162.75 K Boiling point 120.6 °C; 249.0 °F; 393.7 K Refractive index (nD) 1.498 Thermochemistry Heat capacity (C) J K−1 mol−1 Std molarentropy (S298) J K−1 mol−1 Std enthalpy offormation (ΔfH298) kJ mol−1 Std enthalpy ofcombustion (ΔcH298) - kJ mol−1 Hazards GHS labelling: Pictograms Signal word Danger Hazard statements H225, H340, H350, H373 Precautionary statements P201, P210, P308+P313 NFPA 704 (fire diamond) 2 3 0 Flash point 11 °C (52 °F; 284 K) Except where otherwise noted, data are given for materials in their standard state (at 25 °C [77 °F], 100 kPa). Infobox references Chemical compound 1,3-Cycloheptadiene is a highly flammable cycloalkene that occurs as a colorless clear liquid. References This article about a hydrocarbon is a stub. You can help Wikipedia by expanding it.vte Retrieved from " 6 The following pages link to 1,3-Cycloheptadiene External tools: Link count Transclusion count Sorted list Displayed 31 items. View (previous 50 | next 50) (20 | 50 | 100 | 250 | 500)Cyclopentadiene (links | edit) Cyclopropene (links | edit) Cycloalkene (links | edit) Cyclobutadiene (links | edit) Cyclohexene (links | edit) Cyclohexa-1,3-diene (links | edit) Cyclohexa-1,4-diene (links | edit) Cyclooctatetraene (links | edit) Cyclopentene (links | edit) Cyclobutene (links | edit) Cycloheptene (links | edit) Cyclooctadiene (links | edit) 1,5-Cyclooctadiene (links | edit) Trans-Cyclooctene (links | edit) Cis-Cyclooctene (links | edit) Benzene (links | edit) C7H10 (links | edit) Cycloheptadiene (links | edit) Cyclononene (links | edit) Cyclohepta-1,4-diene (links | edit) Cyclononatetraene (links | edit) Cyclodecene (links | edit) Talk:1,3-Cycloheptadiene (links | edit) User:Y-S.Ko/Wikipedia course/Chemistry (links | edit) User:Leyo/InfoCard/0-K (links | edit) User talk:2601:681:4000:72A0:0:0:0:F340 (links | edit) User talk:Siriusbike (links | edit) Wikipedia:WikiProject Chemistry/Lists of pages/Chembox articles (links | edit) Wikipedia:Chemical infobox/Wikipedia:WikiProject Chemistry and Template:Chembox articles (links | edit) Template:Talk:Chem molar mass/articles having chembox enthalpy of formation (links | edit) View (previous 50 | next 50) (20 | 50 | 100 | 250 | 500) Retrieved from " WhatLinksHere/1,3-Cycloheptadiene"