

By. Er. Dharmendra Sir

7974073108, 9584873492

# **DPM CLASSES**

6th to 10th (Math's & Science), 11th & 12th (Physics, Chemistry, Math's)

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**TEST - PAPER (CBSE/NCERT)**

**SOLID STATE**

**SESSION -2024-25**

**CLASS - 12<sup>th</sup>**

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Time : 1 hr : Solid state !

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Q. 1. Why are liquids and gases categorised as fluids ?

Q. 2. Why does table salt, NaCl sometimes appear yellow in colour ?



Q. 3. Why does white  $ZnO(s)$  becomes yellow upon heating ?

Q. 4. Why does the electrical conductivity of semiconductors increase with rise in temperature ?



Q. 5. Explain why does conductivity of germanium crystals increase on doping with gallium ?

Q. 6. Under which situations can an amorphous substance change to crystalline form.

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Q. 7. In a compound, nitrogen atoms (N) make cubic close packed lattice and metal atoms (M) occupy one-third of the tetrahedral voids present. Determine the formula of the compound formed by M and N



Q. 8. Show that in a cubic close packed structure, eight tetrahedral voids are present per unit cell.

Q. 9. How does the doping increase the conductivity of semiconductors?



Q. 10. A sample of ferrous oxide has actual formula  $\text{Fe}_{0.93} \text{O}_{1.00}$ . In this sample, what fraction of metal ions are  $\text{Fe}^{2+}$  ions? What type of non-stoichiometric defect is present in this sample?

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