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TEST - PAPER (CBSE/NCERT)

CHEMICAL KINETICS

SESSION -2024-25

CLASS - 12th

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Add: Block -B, Near Glory Villa, 60 feet Road (Nearest land mark Scindia Statue), Shatabdipuram, Gwalior

Time : 1 hr : Chemical Kinetics :- mm :

Q.1. State a condition under which a bimolecular reaction is kinetically first order reaction.

Q.2. How can you determine the rate law of the following reaction ?

$$2NO(g) + O_2(g) \longrightarrow 2NO_2(g)$$

Q.3. Derive an expression to calculate time required for completion of zero order reaction.

Q.4. For a zero order reaction will the molecularity be equal to zero ? Explain ?

Q.5. What is the probability of reaction with molecularity higher than three very rare ?

Q.6. Why in the redox titration of $KMnO_4$ Vs oxalic acid, we heat oxalic acid solution before starting the titration ?

Q.7. Why molecularity is applicable only for elementry reactions and order is ~~application~~ applicable for elementry as well as complex reactions ?

Q.8. Why can we not determine the order of a reaction by taking into consideration the balanced chemical equation ?

Q.9. Explain the difference b/w instantaneous rate of a reaction and average rate of reaction.

Q.10. With the help of an example explain what is meant by pseudo first order reaction.