

By. Er. Dharmendra Sir

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DPM CLASSES

6th to 10th (Math's & Science), 11th & 12th (Physics, Chemistry, Math's)

TEST - PAPER (CBSE/NCERT)

CHEMICAL REACTIONS AND EQUATIONS

SESSION -2024-25

CLASS - 10th

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Time: 1:00hr Test - Chemistry Reactions
and Equations

mm: 50

Q. 1. Multiple choice questions :-



The above reaction is an example of a -

- (a) Combination reaction (b) double displacement reaction
(c) decomposition reaction (d) displacement reaction.

(ii) What happens when dilute hydrochloric acid is added to iron filings?

- (a) Hydrogen gas and iron chloride are produced.
(b) chlorine gas and iron hydroxide are produced.
(c) No reaction takes place.
(d) Iron salt and water are produced.

(iii) Which among the following is not an oxidizing agent?

- (a) oxygen (b) Conc. Sulphuric acid
(c) chlorine (d) Hydrogen.

(iv) Chemistry rust is -



- (a) Hydrated ferrous oxide (b) Ferric oxide
(c) Hydrated ferric oxide (d) None of these

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(v) What type of reaction is the reaction of calcium oxide with water?

- (a) Endothermic reaction (b) Exothermic reaction
(c) Neutralization reaction (d) Redox reaction.

Q.2. Fill in the Blanks :-

- (i) ----- coloured powder is formed on burning magnesium ribbon in air.
(ii) The formation of water by the reaction of $H_2(g)$ and $O_2(g)$ is a ----- reaction.
(iii) Silver articles turn black on keeping in open due to the formation of -----.
(iv) oily and fatty food items become rancid due to -----.
(v) Those reactions in which heat is ~~not~~ released along with the formation of products are called ----- reactions.

Q.3. Match the columns :-

- | | |
|--------------------------|-------------------------|
| (i) Respiration | (a) Calcium Carbonate |
| (ii) calcium oxide | (b) white powder |
| (iii) magnesium oxide | (c) quick lime |
| (iv) marble | (d) absorption of heat |
| (v) Endothermic reaction | (e) Exothermic reaction |

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Q.3. Why is respiration considered an exothermic reaction?

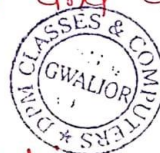
Q.4. A shiny brown coloured element 'X' on heating in air becomes black in colour. Name the element 'X' and the black coloured compound formed.



Q.5. Why do we apply paint on iron articles?

Q.6. Why should a magnesium ribbon be cleaned before it is burnt in air?

Q.7. Define the combination reaction and decomposition reaction.



Q.8. Why do you mean by a precipitation reaction? Explain by giving examples.

Q.9. Explain the following terms with one example each:

(a) Corrosion

(b) Rancidity

Q.10. When sodium sulphate reacts with barium chloride. Which is the name of this type of reaction and write its chemical reaction.

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