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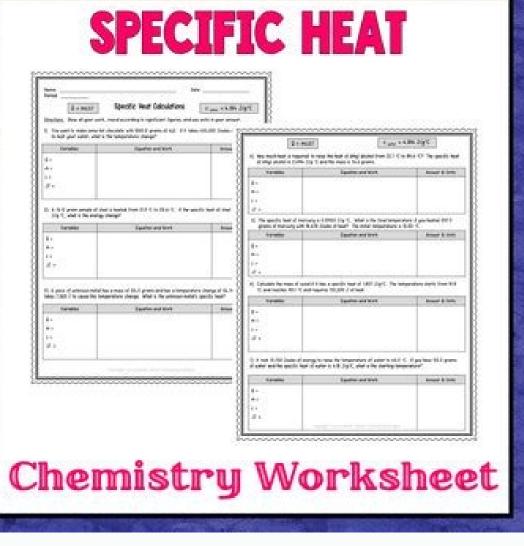
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## Worksheet specific heat

Calculating specific heat worksheet. Specific heat calculations worksheet answers. Specific heat worksheet answers 1-18. Worksheet calculations involving specific heat. Specific latent heat worksheet. Specific heat worksheet #1. Specific heat worksheet (m)(Δt)(csp)=q answer key. Specific heat and calorimetry worksheet. Worksheet specific heat capacity. Specific heat worksheet 1 answer key. Worksheet introduction to specific heat capacities. Specific heat problems worksheet. Specific heat worksheet pdf.

Chemistry worksheet specific heat capacity answers. Specific heat chem worksheet 16-1.

• English as a second language (ESL) • English



Imagine the impact on the southern hemisphere climate.

Specific Heat Worksheet	Name (in ink):
Remember, $\Delta T = (T_{final} - T_{initial})$	ergy, m = mass, and T = temperature  Show all work and proper units.  Ind of the worksheet without units.
	1086.75 joules of heat energy, and its 175°C. Calculate the specific heat capacity of
How many joules of heat are nee aluminum from 22°C to 55°C, if the	eded to raise the temperature of 10.0 g of specific heat of aluminum is 0.90 J/g°C?
	g piece of glass raise if it absorbs 5275 joules y is 0.50 J/g*C? The initial temperature of the
Calculate the heat capacity of a pabsorbs 6.75×10 <sup>4</sup> joules of heat, and	piece of wood if 1500.0 g of the wood d its temperature changes from 32°C to 57°C.
5. 100.0 mL of 4.0°C water is heate heat of water is 4.18 J/g°C, calculate	ed until its temperature is 37°C. If the specific the amount of heat energy needed to cause

Specific Heat Capacities of Various Materials

SUBSTANCE SPECIFIC HEAT CAPACITY (JKG °C)

Aluminum 9.0 x 10 °C

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Brass 3.8 x 10 °Acobol (enthyl) 2.5 x 10 °C

Copper 3.9 x 10 °G

Gisses (crown) 6.7 x 10 °Acobol (methyl) 2.5 x 10 °C

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Lead 1.3 x 10 °Acobol (methyl) 1.1 x 10 °Acobol (methyl) 1.1

Answers 
1.  $T_2 = 47^{\circ}\text{C}$  
2.  $c = 3.8 \times 10^2 \text{ J/kg°C}$  
3. m = 0.37 kg 
4.  $c = 4.4 \times 10^2 \text{ J/kg°C}$ 

Specific Heat Practice Problems

Temperature changes are indicated \ (\ delta t = t\_F - t\_I \) where \ (t\_F \) is the final temperature, O \ (T\_I \) is the original temperature of 15.0 \: \ Text {g} \ Tim 38,7^\ Text {O} \ Text {G} \ Text {G