

SID Study Material



Subject: Science 1

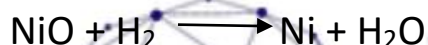
Marks: 25

Chapter: Chemical reactions and equations

DPP-2

I. Answer (5)

1. The conversion of ferrous sulphate to ferric sulphate is..... reaction.
2. Identify from the following reactions the reactants that undergo oxidation and reduction.



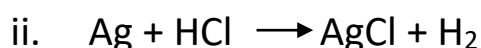
3. is used in the voters-ink.
4. The chemical formula of rust is
5. Name the catalyst used to accelerate the rate of decomposition of hydrogen peroxide.

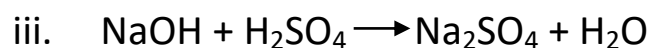
II. Answer any three. (6)

1. Explain the types of reaction with reference to oxygen and hydrogen. Illustrate with examples.
2. Give scientific reasons. While preparing dilute sulphuric acid from concentrated sulphuric acid in the laboratory, the concentrated sulphuric acid is added slowly to water with constant stirring.
3. Write the reactions
 - a. Calcium oxide is mixed with water.
 - b. Ammonia gas reacts with hydrogen chloride gas.
4. Name the various oxidants. How nascent oxygen is liberated from these oxidants.

III. Answer any three (9)

1. Balance the following equations.





2. Define

- i. Unbalanced equation.
- ii. Chemical equation.
- iii. Double displacement reaction,

3. Explain redox reactions with two examples.

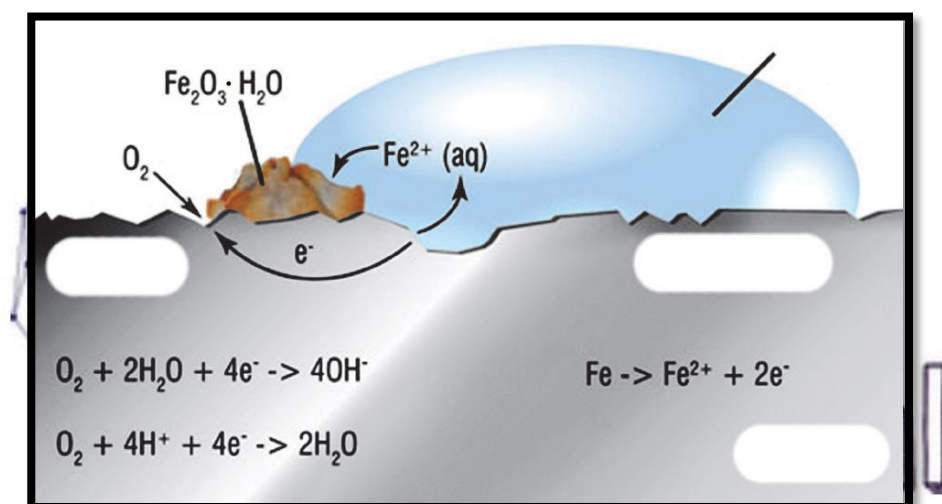
4. Write five conventions used to write chemical equation.

IV. Answer any one

(5)

1. Explain the factors affecting rate of reaction.

2. Observe the following picture and write down the chemical reactions with explanation by completing the diagram.



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