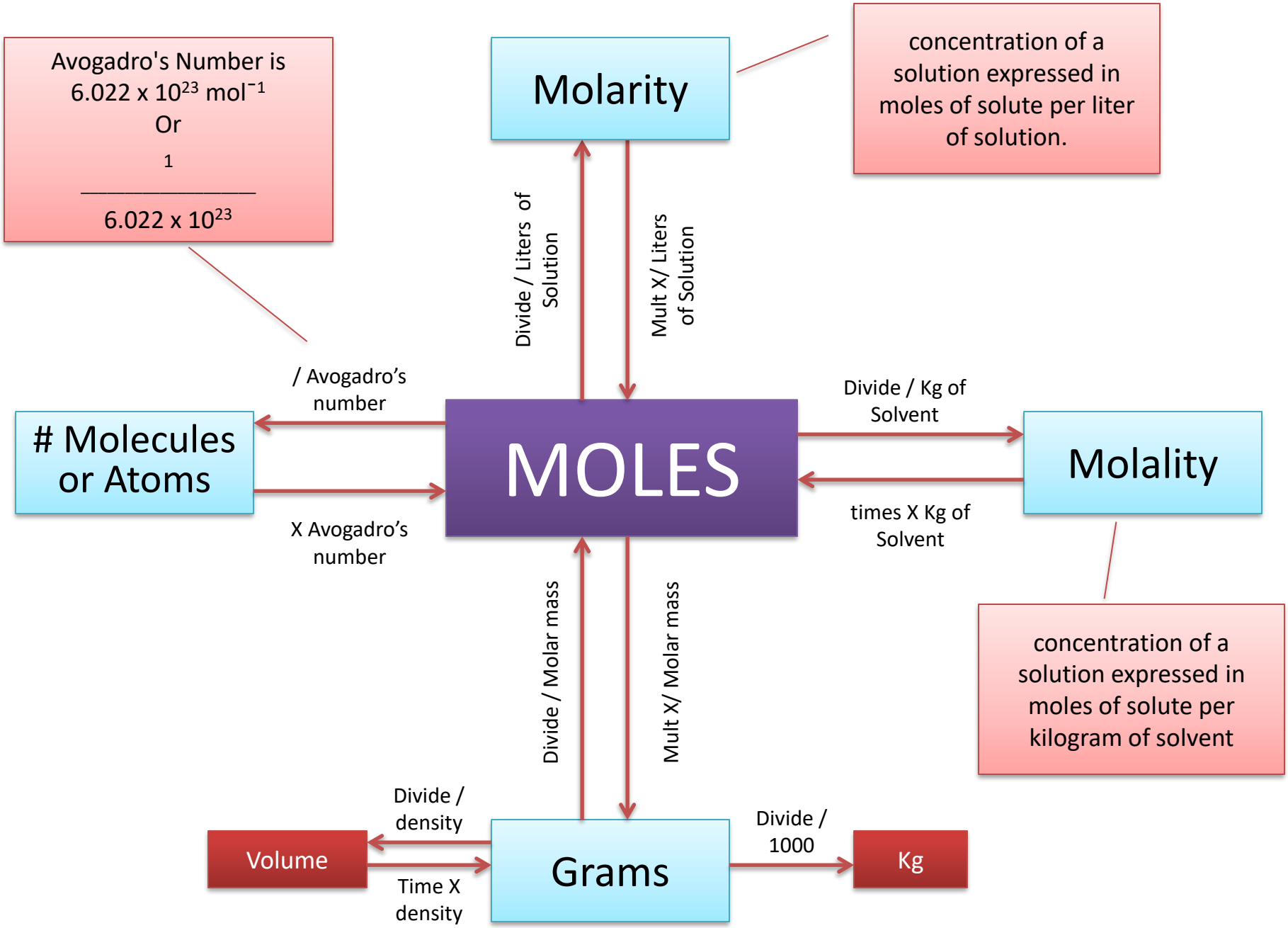
Avogadro's Number is
6.022 x 10^23 mol^-1
Or
1
6.022 x 10^23
Molarity
concentration of a solution expressed in moles of solute per liter of solution.
Divide / Liters of Solution
Mult X/ Liters of Solution
/ Avogadro's number
Molecules or Atoms
X Avogadro's number
MOLES
Divide / Kg of Solvent
times X Kg of Solvent
Molality
concentration of a solution expressed in moles of solute per kilogram of solvent
Divide / Molar mass
Mult X/ Molar mass
Divide / density
Volume
Time X density
Grams
Divide / 1000
Kg