



# U. S. OSTWAL INTERNATIONAL SCHOOL

OSTWAL WONDER CITY, BOISAR (E)

TERM – 1 (2023-24)

Name : \_\_\_\_\_

Date: \_\_\_\_\_

Roll no. \_\_\_\_\_

Marks: 60

Grade : VIII

Time : 2 hr 30 min

Subject : Chemistry

## General Instructions:

- All questions are compulsory
- Do not write the questions. Directly write the answers.
- Write the paper neatly. Reading Time: 15 minutes.

### Q.1) Fill in the blanks.

(06)

- 1) The \_\_\_\_\_ is positively charged Central part of the atom.
- 2) \_\_\_\_\_ shell is the outermost shell of an atom
- 3) A \_\_\_\_\_ bond is formed by the sharing of electrons.
- 4) An \_\_\_\_\_ is the simplest form of a pure substance which is made up of only one kind of atom.
- 5) \_\_\_\_\_ refers to the combining capacity of an element.
- 6) Decomposition of water into hydrogen and oxygen occurs in the presence of \_\_\_\_\_

### Q.2) Choose the correct answer.

(06).

- 1) An atom consists of \_\_\_\_\_  
a) Electrons    b) proton    c) neutrons    d) all of these
- 2) \_\_\_\_\_ suggested that electrons possess specific amount of energy and are arranged in a succession of energy levels  
a) J.J. Thomson    b) James Chadwick    c) Niels Bohr    e) E. Rutherford
- 3) which of the following elements does not exhibit variable valency?  
a) Copper    b) iron    c) phosphorus    d) lead
- 4) The symbol Na of sodium is derived from its Latin name \_\_\_\_\_.  
a) Nalium    b) natium    c) natrium    d) none of these
- 5) which of the following elements is represented by the symbol Ag?  
a) Gold    b) Silver    c) Potassium    d) Copper
- 6) which of the following elements is tetravalent?.  
a) Silicon.    b) Lithium.    c) Boron    d) Oxygen

### Q. 3) Write the atomic number and mass number of the following elements.

(06)

Element	Atomic number	Mass number
Lithium		
Carbon		
Fluorine		
Magnesium		
Sulphur		
Potassium		

**Q 4 Match the columns****(06)**

## Column A.

1. Anode rays
2. J.J.Thomson.
3. Magnesium.
4. James Chadwick.
5. Chlorine.
6. Phosphorus

## Column B

- a) discovery of neutron
- b) monovalent element
- c) trivalent element
- d) consist of protons
- e) divalent element
- f) plum pudding model

**Q.5) Answer in shorts .(Any 5)****(10)**

- 1) Define relative atomic mass.
- 2) What are isotopes? Give examples.
- 3) How is an anion formed?
- 4) Why do atoms of elements combine or react?
- 5) Balance the following equations



- 6) Write the molecular formulae of some common compound.

a) Hydrogen chloride

b) Sodium hydroxide

**Q.6) Show the given text in the form of flowcharts. (Any 2)****(06)**

- 1) Types of chemical bond
- 2) Types of chemical equations
- 3) Types of ions

**Q.7) Differentiate between the following. (Any 2)****(06)**

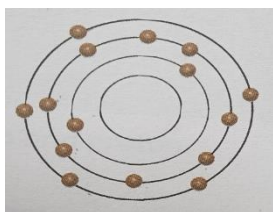
1. Elements and compound
2. Positive valency and negative valency
3. Electrovalent bond and chemical bond

**Q. 8 Give reasons for the following statements . (Any 2)****(04)**

- 1) Cathode rays consist of negatively charged particles.
- 2) An atom is electrically neutral.
- 3) Photosynthesis is considered as an endothermic reaction.

**Q. 9 Write the Latin names of Elements.****(05)**

1. Sodium \_\_\_\_\_
2. Potassium \_\_\_\_\_
3. Iron \_\_\_\_\_
4. Copper \_\_\_\_\_
5. Gold \_\_\_\_\_

**Q.10 Observe the given image carefully and answer the following questions.****(05)**

1. Name the element represented through the image.
2. What is the atomic number and mass number of the element ?
3. Write the number of electrons, protons and neutrons in the element.

**\*ALL THE BEST\***

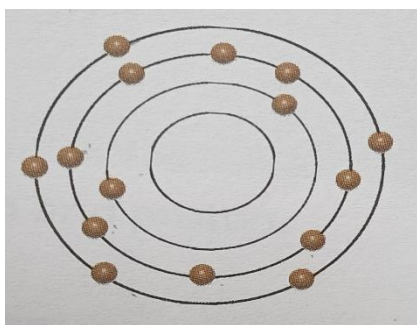
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**(05)**

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**Q.10 Observe the given image carefully and answer the following questions.**

**(05)**



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