



U. S. OSTWAL INTERNATIONAL SCHOOL

OSTWAL WONDER CITY, BOISAR (E)

TERM – 1 (2023-24)

Name : _____

Date:

Roll no. _____

Max. marks : 60

Grade : VII

Time : 2hr30min

Subject : Chemistry

General Instructions:

- All questions are compulsory
- Do not write the questions. Directly write the answers.
- Write the paper neatly.
- 10 min for reading

Q.1) Answer in one word / a few words. (06)

- 1) The simplest form of a pure substance made up of only one kind of atom. _____
- 2) The number of atoms contained in a molecule of an element _____
- 3) A homogeneous mixture of solvent and solute . _____
- 4) The smallest indivisible particle of matter that cannot exist independently. _____
- 5) Representation of bromide ion . _____
- 6) Chemical formula of sugar. _____

Q.2) Choose the correct answer. (06)

1) Which of the following metals is brittle and crumbles into powder ?

- a) Mercury
- b) Zinc
- c) Potassium
- d) Calcium

2) What is the symbol of sodium?.

- a) S
- b) So
- c) Na
- d) Si

3) What is the common name of sodium bicarbonate?

- a) Common salt
- b) Washing soda
- c) Vinegar
- d) Baking soda

4) Which of the following represents a water molecule?

- a) HO
- b) H₂O
- c) HO₃

d) H₂O

5) Which of the following elements possesses variable valency?

- a) Fluorine
- b) Sodium
- c) Iron
- d) Silicon

6) How many groups are there in a periodic table ?

- a) 13
- b) 15
- c) 17
- d) 18

Q. 3 Identify the error (s) in the given statements. (06)

1. Atoms have more than one kind of particles combined together in a fixed ratio by mass.
2. Sugar solution is a heterogenous mixture. _____
3. Oil and water are miscible liquids _____
4. Ozone is a polyatomic molecule. _____
5. A negatively charged ion is called cation. _____
6. Chlorine ion is represented as Ca²⁺ _____

Q. 4 Match the columns. (06)

Column A

1. Stones and water
2. Salt in water
3. Water in flour
4. Vinegar in water
5. Fog
6. Aerated drink

Column B

- a) Liquid – gas (heterogeneous) mixture
- b) Liquid – liquid (heterogeneous) mixture
- c) Solid – liquid (homogeneous) mixture
- d) Gas – liquid (homogeneous) mixture
- e) Solid – solid (heterogeneous) mixture
- f) Liquid – solid (heterogeneous) mixture

Q.5) Answer in shorts .

(Any 6)

(12)

- 1) What is atomicity?
- 2) What is sublimation?
- 3) What is molecular formula?
- 4) What are noble gases? Give examples.
- 5) What is an atom ? Give examples.
- 6) Name the types of molecules.
- 7) How is a cation different from an anion

Q.6 Show the given text in the form of flowcharts. (06)

(Any 2)

- 1) Types of mixtures
- 2) Classification of elements on the basis of valency
- 3) Atoms , molecules and radicals

Q.7 Differentiate between the following.

(Any 2)

(06)

1. Atomicity and valency
2. Metals and non- metals
3. Atoms and molecules

Q. 8 Give reasons for the following statements .

(10)

(Any 5)

- 1) Elements and compounds are pure substances.
- 2) The composition of a mixture is variable.
- 3) Separation of various components of a mixture is vital.
- 4) An atom combines with other atoms or molecules.
- 5) An ion is a charged particle.
- 6) An atom cannot be further divided.

Q. 9 Write the chemical formula of the following compounds.

(02)

1. Common salt
2. Hydrochloric acid

ALL THE BEST

